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Syllabus and Scheme of Examination for Chemistry (Honors) under Choice Based Credit System (CBCS) Curricula

UNIVERSITY OF GOUR BANGA, MALDA

Draft Scheme/Broad Guidelines of UG CBCS Curriculum

Meaning of the Terms and Abbreviations

- ***Discipline Core (DC) Course:*** A course that should compulsorily be studied by a candidate as a core requirement of the programme.
- ***Elective Course:*** A Course which may be very specific or specialized or advanced or supportive to the discipline/ subject of study. These are of three types:
  - ***A) Discipline Specific Elective (DSE) Course:*** A course, which may be offered by the main discipline/subject of study.
  - ***B) Generic Elective (GE) Course:*** An elective course, chosen from related or unrelated discipline/subject of study.
  - ***C) Dissertation/Project (D.P):*** An elective course designed to acquire special/advanced knowledge with an advisory support by a teacher/faculty member.
- ***Ability Enhancement Compulsory (AEC) Course:*** The course designed for knowledge enhancement consisting of Environmental studies, English Communication/MIL.
- ***Skill Enhancement Course (SEC):*** These courses are designed to provide value-based and/or skill-based knowledge relating to the main discipline.
- ***MCQ:*** Multiple Choice Questions
- ***MIL:*** Modern Indian Language
- ***DQ:*** Descriptive questions
- ***IA:*** Internal Assessment (By the respective Department of the Colleges).
- ***Teaching terms:*** Lectures **(L) /** Tutorials (**Tu) /** Theory **(Th**) / Practical **(Pr) /** Dissertation/Project Work **(DP) /** Viva **(V) /** Seminar **(S) /** Term Paper **(TP).**
- ***ESE:*** End Semester Examination
- ***TBD:*** To Be Decieted

Semester wise course structure under CBCS for B.A/B.Sc./B.Com (Hons.) Courses

Academic Semesters	C O U R S E S					Credits	Marks
	Discipline Core (DC)	Discipline Specific Elective (DSE)	Generic Elective (GE)	Ability Enhancement Compulsory (AEC)	Skill Enhancement (SEC)		
SEM-I	DC1(6) DC2(6)	-	GE-1 (6)	ENVS (2)	-	20	200
SEM-II	DC3(6) DC4(6)	-	GE-2 (6)	Communicative English/Communicative Bengali/MIL (2)	-	20	200
SEM-III	DC5(6) DC6(6) DC7(6)	-	GE-3 (6)	-	-	24	200
SEM-IV	DC8(6) DC9(6) DC10(6)	-	EEE EG E-4 (6)	-	-	24	200
SEM-V	DC11(6) DC12(6)	DSE-1 (6) DSE-2 (6)	-	-	SE-1 (2)	26	250
SEM-VI	DC13(6) DC14(6)	DSE-3 (6) DSE-4/ DP-4(6)	-	-	SE-2 (2)	26	250
Total	-	-	-	-	-	**140**	**1300**

Notes:

1. **Each course is of 50 marks**
2. **DC, DSE ,GE & DP:** Each course is of 6 credits

   Non-practical course: Theory - 5 credits + Tutorial- 1 credit

   Practical based course: Theory - 4 credits + Practical- 2 credits
3. **Credit =** 1 hour duration of teaching (lecture/tutorial) or 2 hour duration of practical period.
4. **AEC and SEC:** Each course is of 2 credits having no tutorial/practical period
5. **Optional:** Dissertation/ Project Work in place of one DSE in 6th semester.
6. An honours student has to study two disciplines as GE. Each discipline will provide 4 papers as GE courses for SEM I to SEM IV. An Hons. Candidate has to study four GE papers from two disciplines other than concerned Hons. discipline.
7. For each of the DSE paper, a student has to select one paper out of two options provided.
8. A student other than Chemistry honours will take GE (General Elective) papers from the syllabus for Chemistry (General) as mentioned under Displine Core (DC).

Marks and Question type Distributions for Hons. Courses

No. of courses	Total credit	Total marks	Division of marks of each course					Marks for question type
			Full marks of each course	Internal	End semester examination		MCQ	Written
					Theoretical	Practical		
14 DC	14x6 =84	14x50=700	50 (practical based)	10	25	15	Nil	40
04 DSE	04x6=24	4x50=200	50 (practical based)	10	25	15	Nil	40
04 GE	04x6=24	4x50=200	50 (practical based)	10	25	15	Nil	40
02SE	02x2=04	2x50=100	50	10	40	Nil	Nil	40
AEC-1 (ENVS)	01x2=02	1x50=50	50	10 (Project to be internally assessed)	40	Nil	Nil	40
AEC-2 (Communicative Bengali/English)	01x2=02	1x50=50	50	10	40	Nil	Nil	40
Grand Total	**140**	**1300**	-	-	-	-	-	-

Distribution of credits for B.Sc (Hons.) in Chemistry Courses

Course type and credit	Number of Courses X credit (With practical) [Theory:04+Practical:02]		Number of Courses X 02 Credit (Without practical)	
	Theory	**Practical**	**Theory**	**Tutorial**
Discipline Course (DC) (6)	14x4=56	14x2=28	NA	NA
Generic Elective (GE) (6)	4x4=16	4x2=8	NA	NA
Discipline Specific Elective (DSE) (6)	4x4=16	4x2=8	NA	NA
Ability Enhancement Course (AEC) (2)	NA	NA	2x2=4	00
Skill Enhancement (SE) (2)	NA	NA	2x2=4	00
Courses Credit	88	44	8	00
Total credit	**140**			

Detail Curriculum (Chemistry Honors, Semester wise)

Semester	Course	Paper Code	Broad area	Credit
I	Core Course-1 (Theory) (**DCT1**)	CEMHT-1	Organic-I	4
	Core Course-1 (Practical) (**DCP1**)	CEMHP-1	Organic-I	2
	Core Course-2 (Theory) (**DCT2**)	CEMHT-2	Physical - I	4
	Core Course-2 (Practical)(**DCP2**)	CEMHP-2	Physical - I	2
	Generic Elective-1 (Theory) *	TBD	TBD	4
	Generic Elective-1 (Practical)	TBD	TBD	2
	Ability Enhancement Compulsory Course-11	TBD	*English communication/ Environmental Science	2
II	Core Course-3 (Theory)(**DCT3**)	CEMHT-3	Inorganic-I	4
	Core Course-3(Practical)(**DCP3**)	CEMHP-3	Inorganic-I	2
	Core Course-4 (Theory)(**DCT4**)	CEMHT-4	Organic - II	4
	Core Course-4 (Practical) (**DCP4**)	CEMHP-4	Organic - II	2
	Generic Elective-2 (Theory)	TBD	TBD	4
	Generic Elective-2 (Practical)	TBD	TBD	2
	Ability Enhancement Compulsory Course-2	TBD	*English Communication/ Environmental Science	2
III	Core Course-5 (Theory)(**DCT5**)	CEMHT-5	Physical – II	4
	Core Course-5 (Practical)(**DCP5**)	CEMHP-5	Physical – II	2
	Core Course-6 (Theory)(**DCT6**)	CEMHT-6	Inorganic - II	4
	Core Course-6 (Practical)(**DCP6**)	CEMHP-6	Inorganic - II	2
	Core Course-7 (Theory)(**DCT7**)	CEMHT-7	Organic-III	4
	Core Course-7 (Practical)(**DCP**)	CEMHP-7	Organic-III	2
	Generic Elective-3 (Theory)	TBD	TBD	4
	Generic Elective-3 (Practical)	TBD	TBD	2
IV	Core Course-8 (Theory)(**DCT8**)	CEMHT-8	Physical – III	4
	Core Course-8 (Practical)(**DCP8**)	CEMHP-8	Physical – III	2
	Core Course-9 (Theory)(**DCT9**)	CEMHT-9	Inorganic - III	4
	Core Course-9 (Practical)(**DCP9**)	CEMHP-9	Inorganic - III	2
	Core Course-10(Theory)(**DCT10**)	CEMHT-10	Organic-IV	4
	Core Course-10 (Practical)(**DCP10**)	CEMHP-10	Organic-IV	2
	Generic Elective-4 (Theory)	TBD	TBD	4
	Generic Elective-4 (Practical)	TBD	TBD	2
V	Core Course-11 (Theory)(**DCT11**)	CEMHT-11	Inorganic - IV	4
	Core Course-11 (Practical)(**DCP11**)	CEMHP-11	Inorganic - IV	2
	Core Course-12 (Theory)(**DCT12**)	CEMHT-12	Organic - V	4
	Core Course-12 (Practical)(**DCP12**)	CEMHP-12	Organic - V	2
	Discipline Specific Elective- 1	CEMHTDSE-1A	Advanced Physical	4

	(Theory) (Compulsory) (**DSET1**)		Chemistry	
	Discipline Specific Elective-1 (Practical) (Compulsory) (**DSEP1**)	CEMHPDSE- 1A	Advanced Physical Chemistry	2
	Discipline Specific Elective-1 (Theory) (**DSET2**)	CEMHTDSE-2A	Analytical Methods in Chemistry	4
		CEMHTDSE-2B	Polymer Chemistry	
	Discipline Specific Elective-1 (Practical) (**DSEP2**)	CEMHPDSE- 2A	Analytical Methods in Chemistry	2
		CEMHPDSE- 2B	Polymer Chemistry	
	Skill Enhancement Course – 2 (**SE-1**) (Anyone from this group)	CEMHSE-1A	IT skills for chemists	2
		CEMHSE-1B	Basic Analytical Chemistry	
VI	Core Course-13 (Theory)(**DCT13**)	CEMHT-13	Inorganic - V	4
	Core Course-13(Practical)(**DCP13**)	CEMHP-13	Inorganic - V	2
	Core Course-14(Theory)(**DCT14**)	CEMHT-14	Physical-IV	4
	Core Course-14(Practical)(**DCP14**)	CEMHP-14	Physical-IV	2
	Discipline Specific Elective-3 (Theory) (**DSET3**) (Any One)	CEMHTDSE- 3A	Instrumental methods of Chemical Analysis	4
		CEMHTDSE-3B	Inorganic materials of Industrial importance	4
	Discipline Specific Elective-2 (Practical) (**DSEP3**) (Any One)	CEMHPDSE-3A	Instrumental methods of Chemical Analysis	2
		CEMHPDSE-3B	Inorganic materials of Industrial importance	2
	Discipline Specific Elective-4 (Theory) (**DSET4**)**	CEMHTDSE-4	Green Chemistry	4
	Discipline Specific Elective- 4 (Practical)(**DSEP4**)**	CEMHPDSE-4	Green Chemistry	2
	Discipline Specific Elective- 4 **(DSE-4)****	DSEPROJ- 5	Project	6
	Skill Enhancement Course – 2 (**SE-2**) (Anyone from this group)	CEMHSE-2A	Pharmaceutical Chemistry	2
		CEMHSE-2B	Analytical Clinical Biochemistry	

** Dissertation followed by power point presentation (credit point 6) may be alternative option of taking "Green Chemistry" as Discipline Specific Elective-4 (DSE-4) and paper code is "DSEPROJ-5"

Semester - I

CEMHT-1	Organic Chemistry – I	4 Credit

Theory: Bonding and Physical Properties, General Treatment of Reaction Mechanism-I, Stereochemistry- I

1. Bonding and Physical Properties: (17L)

Valence Bond Theory:

Concept of hybridisation, shapes of molecules, resonance (including hyperconjugation); calculation of formal charges and double bond equivalent (DBE); orbital pictures of bonding (sp^3, sp^2, sp: C-C, C-N & C-O systems and s-cis and s-trans geometry for suitable cases).

Electronic displacements:

Inductive effect, field effect, mesomeric effect, resonance energy; bond polarization and bond polarizability; electromeric effect; steric effect, steric inhibition of resonance.

MO theory:

Qualitative idea about molecular orbitals, bonding and antibonding interactions, idea about σ, σ*, π, π*, n – MOs; basic idea about Frontier MOs (FMO); concept of HOMO, LUMO and SOMO; interpretation of chemical reactivity in terms of FMO interactions; sketch and energy levels of π MOs of i) acyclic p orbital system (C=C, conjugated diene, triene, allyl and pentadienyl systems), ii) cyclic p orbital system (neutral systems: [4], [6]-annulenes; charged systems: 3-, 4-, 5- membered ring systems); Hückel's rules for aromaticity up to [10]-annulene (including mononuclear heterocyclic compounds up to 6-membered ring); concept of antiaromaticity and homoaromaticity; non-aromatic molecules; Frost diagram; elementary idea about α and β; measurement of delocalization energies in terms of β for buta-1,3-diene, cyclobutadiene, hexa-1,3,5-triene and benzene.

Physical Properties:

Influence of hybridization on bond properties: bond dissociation energy (BDE) and bond energy; bond distances, bond angles; concept of bond angle strain (Baeyer's strain theory); melting point/boiling point and solubility of common organic compounds in terms of covalent & non-covalent intermolecular forces; polarity of molecules and dipole moments; relative stabilities of isomeric hydrocarbons in terms of heat of hydrogenation, heat of combustion and heat of formation.

2. General Treatment of Reaction Mechanism – I: (25L)

Mechanistic Classification:

Ionic, radical and pericyclic (definition and example); reaction type: addition, elimination and substitution reactions (definition and example); nature of bond cleavage and bond formation: homolytic and heterolytic bond fission, homogenic and heterogenic bond formation; curly arrow rules in representation of mechanistic steps; reagent type: electrophiles and nucleophiles (elementary idea); electrophilicity and nucleophilicity in terms of FMO approach.

Reactive intermediates:

Carbocations (carbenium and carbonium ions), carbanions, carbon radicals, carbenes, benzyne, nitrenes: generation and stability, structure using orbital picture and electrophilic/nucleophilic behavior of reactive intermediates (elementary idea).

3. Stereochemistry-I: (18L)

Bonding geometries of carbon compounds and representation of molecules: Tetrahedral nature of carbon and concept of asymmetry; Fischer, sawhorse, flying-wedge and Newman projection formulae and their inter translations.

Concept of chirality and symmetry:

Symmetry elements and point groups (C_v,D_h,C_{nh}, C_{nv}, C_n, D_{nh}, D_{nd}, D_n, S_n (C_s, Ci); molecular chirality and centre of chirality; asymmetric and dissymmetric molecules; enantiomers and diastereomers; concept of epimers; concept of stereogenicity, chirotopicity and pseudoasymmetry; chiral centres and number of stereoisomerism: systems involving 1/2/3-chiral centre(s) (AA, AB, ABA and ABC types).

Relative and absolute configuration:

D/L and R/S descriptors; erythro/threo and meso nomenclature of compounds; syn/anti nomenclatures for aldols; E/Z descriptors for C=C, conjugated diene, triene, C=N and N=N systems; combination of R/S- and E/Z- isomerisms.

Optical activity of chiral compounds:

Optical rotation, specific rotation and molar rotation; racemic compounds, racemisation (through cationic, anionic, radical intermediates and through reversible formation of stable achiral intermediates); resolution of acids, bases and alcohols via diastereomeric salt formation; optical purity and enantiomeric excess; invertomerism of chiral trialkylamines.

Reference Books:

1.Clayden, J., Greeves, N. & Warren, S. Organic Chemistry, Second edition, Oxford University Press, 2012. 2. Smith, J. G. Organic Chemistry, Tata McGraw-Hill Publishing Company Limited. 3. Nasipuri, D. Stereochemistry of Organic Compounds, Wiley Eastern Limited. 4. Morrison, R. N. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 5. Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd., (Pearson Education). 6. Fleming, I. Molecular Orbitals and Organic Chemical Reactions, Reference/Student Edition, Wiley, 2009. 7. Eames, J., Peach, J. M. Stereochemistry at a Glance, Blackwell Publishing, 2003. 8. Robinson, M. J., Stereochemistry, Oxford Chemistry Primer, Oxford University Press, 2005.

CEMHP-1	**Organic Chemistry – I**	**2 Credits**

Practical

1. **Separation:**
   Based upon solubility, by using common laboratory reagents like water (cold, hot), dil. HCl, dil. NaOH, dil. $NaHCO_3$, etc., of components of a binary solid mixture; purification of any one of the separated components by crystallization and determination of its melting point. The composition of the mixture may be of the following types: Benzoic acid/p-Toluidine; p-Nitrobenzoic acid/p- Aminobenzoic acid; p-Nitrotolune/p-Anisidine; etc.

2. **Determination of boiling point:**
   Determination of boiling point of common organic liquid compounds e.g., ethanol, cyclohexane, chloroform, ethyl methyl ketone, cyclohexanone, acetylacetone, anisole, crotonaldehyde, mesityl oxide, etc. [Boiling point of the chosen organic compounds should preferably be less than 160 °C]

3. **Identification of a Pure Organic Compound by chemical test(s): Solid compounds:**
   oxalic acid, tartaric acid, citric acid, succinic acid, resorcinol, urea, glucose, cane sugar, benzoic acid and salicylic acid.
   Liquid Compounds:
   formic acid, acetic acid, methyl alcohol, ethyl alcohol, acetone, aniline, dimethylaniline, benzaldehyde, chloroform and nitrobenzene.

Reference Books:

1. Bhattacharyya, R. C, A Manual of Practical Chemistry. 2. Vogel, A. I. Elementary Practical Organic Chemistry, Part 2: Qualitative Organic Analysis, CBS Publishers and Distributors. 3. Mann, F. G. & Saunders, B. C. Practical Organic Chemistry, Pearson Education (2009). 4. Furniss, B.S., Hannaford, A.J., Smith, P.W.G., Tatchell, A.R. Practical Organic Chemistry,5th Ed., Pearson (2012)

CEMHT-2	Physical Chemistry – I	4 Credit

Theory: Kinetic Theory and Gaseous State, Chemical Thermodynamics, Chemical Kinetics

1. Kinetic Theory and Gaseous State: (18L)

Kinetic Theory of Gases:

Concept of pressure and temperature; Collision of gas molecules; Collision diameter; Collision number and mean free path; Frequency of binary collisions (similar and different molecules).Wall Collision and rate of effusion.

Maxwell's Distribution of Speed and Energy:

Nature of distribution of velocities, Maxwell's distribution of speeds in one, two and three dimensions; Kinetic energy distribution in one, two and three dimensions, calculations of average, root mean square and most probable values in each case; Calculation of number of molecules having energy $\geq \varepsilon$, Principle of equipartition of energy and its application to calculate the classical limit of molar heat capacity of gases.

Real gas and Virial Equation:

Deviation of gases from ideal behavior; compressibility factor; Boyle temperature; Andrew's and Amagat's plots; van der Waals equation and its features; its derivation and application in explaining real gas behaviour, other equations of state (Berthelot, Dietrici); Existence of critical state, Critical constants in terms of van der Waals constants; Law of corresponding states; virial equation of state; van der Waals equation expressed in virial form and significance of second virial coefficient; Intermolecular forces (Debye, Keesom and London interactions; Lennard - Jones potential - elementary idea

2. Chemical Thermodynamics: (24L)

Zeroth and 1st law of Thermodynamics:

Intensive and extensive variables; state and path functions; isolated, closed and open systems; zeroth law of thermodynamics; Concept of heat, work, internal energy and statement of first law; enthalpy, H; relation between heat capacities, calculations of q, w, U and H for reversible, irreversible and free expansion of gases (ideal and van der Waals) under isothermal and adiabatic conditions; Joule's experiment and its consequence.

Thermochemistry:

Standard states; Heats of reaction; enthalpy of formation of molecules and ions and enthalpy of combustion and its applications; Laws of thermochemistry; bond energy, bond dissociation energy and resonance energy from thermochemical data, Kirchhoff's equations and effect of pressure on enthalpy of reactions. Adiabatic Flame Temperature, Explosion Temperature

Second Law:

Need for a Second law: statement of the second law of thermodynamics; Concept of heat reservoirs and heat engines; Carnot cycle; Physical concept of Entropy; Carnot engine and refrigerator; Kelvin –Planck and Clausius statements and equivalence of the two statements with entropic formulation; Carnot's theorem; Values of §dQ/T and Clausius inequality; Entropy change of systems and surroundings for various processes and transformations; Entropy and unavailable work; Auxiliary state functions (G and A) and their variation with T, P and V. Criteria for spontaneity and equilibrium.

Thermodynamic relation: Maxwell's relations; Gibbs-Helmholtz equation, Joule-Thomson experiment and its consequences; inversion temperature; Joule- Thomson coefficient for a van der Waals gas; General heat capacity relations.

3. Chemical Kinetics: (18L)

Rate Law, Order and Molecularity:

Introduction of rate law, Extent of reaction; rate constants, order; Forms of rates of First, second and nth order reactions; Pseudo first order reactions (example using acid catalyzed hydrolysis of methyl acetate); Determination of order of a reaction by half -life and differential method; Opposing reactions, consecutive reactions and parallel reactions (with explanation of kinetic and thermodynamic control of products; all steps first order).

Role of Temperature and Theories of Reaction Rate:

Temperature dependence of rate constant; Arrhenius equation, energy of activation; Rate-determining step and steady-state approximation – explanation with suitable examples; Collision theory; Lindemann theory of unimolecular reaction; outline of Transition State theory (classical treatment).

Homogeneous Catalysis:

Homogeneous catalysis with reference to acid-base catalysis; Primary kinetic salt effect; Enzyme catalysis; Michaelis-Menten equation, Lineweaver-Burk plot, turn-over number. Autocatalysis; periodic reactions

Reference Books:

1.Atkins, P. W. & Paula, J. de Atkins' Physical Chemistry, Oxford University Press. 2. Castellan, G. W. Physical Chemistry, Narosa.3. McQuarrie, D. A. & Simons, J. D. Physical Chemistry: A Molecular Approach, Viva Press.4. Engel, T. & Reid, P. Physical Chemistry, Pearson. 5. Levine, I. N. Physical Chemistry, Tata McGraw-Hill. 6. Maron,S. & Prutton Physical Chemistry. 7. Ball, D. W. Physical Chemistry, Thomson Press. 8. Mortimer, R. G. Physical Chemistry, Elsevier. 9. Laidler, K. J. Chemical Kinetics, Pearson. 10. Glasstone, S. & Lewis, G.N. Elements of Physical Chemistry. 11. Rakshit, P.C., Physical Chemistry Sarat Book House. 12. Zemansky, M. W. & Dittman, R.H. Heat and

Thermodynamics, Tata-McGraw-Hill. 13. Rastogi, R. P. & Misra, R.R. An Introduction to Chemical Thermodynamics, Vikas. 14. Clauze & Rosenberg, Chemical Thermodynamics.

CEMHP-2	Physical Chemistry – I	2 Credit

Practical

1) Determination of pH of unknown solution (buffer), by colour matching method.
2) Study of kinetics of acid-catalyzed hydrolysis of methyl acetate.
3) Determination of heat of solute ion of oxalic acid from solubility measurement.
4) Study of kinetics of decomposition of H_2O_2
5) Determination of heat of neutralization of a strong acid by a strong base.

Reference Books:

1.Viswanathan, B., Raghavan, P.S. Practical Physical Chemistry Viva Books (2009).2.Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson. 3. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman (2007). 4. Palit, S.R., De, S. K. Practical Physical Chemistry Science Book Agency. 5. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N., University of Calcutta. 6. Levitt, B. P. edited Findlay's Practical Physical Chemistry Longman Group Ltd. 7. Gurtu, J. N., Kapoor, R., Advanced Experimental ChemistryS. Chand & Co. Ltd.Ltd.

Semester-II

CEMHT-3	Inorganic Chemistry-I	4 Credit

Theory: Extra nuclear Structure of atom, Chemical periodicity, Acid-Base reactions, Redox Reactions and precipitation reactions

1. Extra nuclear Structure of atom: (15L)

Bohr's theory, its limitations and atomic spectrum of hydrogen atom; Sommerfeld's Theory. Wave mechanics: de Broglie equation, Heisenberg's Uncertainty Principle and its significance, Schrödinger's wave equation, significance of ψ and ψ^2. Quantum numbers and their significance. Radial and angular wave functions for hydrogen atom. Radial and angular distribution curves. Shapes of s, p, d and f orbitals. Pauli's Exclusion Principle, Hund's rules and multiplicity, Exchange energy, Aufbau principle and its limitations, Ground state Term symbols of atoms and ions foratomic number upto 30.

2. Chemical Periodicity: (15L)

Modern IUPAC Periodic table, Effective nuclear charge, screening effects and penetration, Slater's rules, atomic radii, ionic radii (Pauling's univalent), covalent radii, lanthanide contraction. Ionization potential, electron affinity and electronegativity (Pauling's, Mulliken's and AllredRochow's scales) and

factors influencing these properties, group electronegativities. Group trends and periodic trends in these properties in respect of s-, p- and d-block elements. Secondary periodicity, Relativistic Effect, Inert pair effect.

3. Acid-Base Reactions: (15L)

Acid-Base concept: Arrhenius concept, theory of solvent system (in H_2O, NH_3, SO_2 and HF), Bronsted-Lowry's concept, relative strength of acids, Pauling's rules. Lux-Flood concept, Lewis concept, group characteristics of Lewis acids, solvent levelling and differentiating effects. Thermodynamic acidity parameters, Drago-Wayland equation. Superacids, Gas phase acidity and proton affinity; HSAB principle. Acid-base equilibria in aqueous solution (Proton transfer equilibria in water), pH, buffer. Acid-base neutralization curves; indicator, choice of indicators.

4. Redox Reactions and Precipitation Reactions: (15L)

Ion-electron method of balancing equation of redox reaction. Elementary idea on standard redox potentials with sign conventions, Nernst equation (without derivation). Influence of complex formation, precipitation and change of pH on redox potentials; formal potential. Feasibility of a redox titration, redox potential at the equivalence point, redox indicators. Redox potential diagram (Latimer and Frost diagrams) of common elements and their applications. Disproportionation and comproportionation reactions (typical examples). Solubility product principle, common ion effect and their applications to the precipitation and separation of common metallic ions as hydroxides, sulfides, phosphates, carbonates, sulfates and halides.

Reference Books:

1.Lee, J. D. Concise Inorganic Chemistry ELBS, 1991. 2. Douglas, B.E. and McDaniel,D.H. Concepts & Models of Inorganic Chemistry Oxford, 1970. 3. Day, M.C. and Selbin, J. Theoretical Inorganic Chemistry, ACS Publications, 1962. 4. Atkin, P. Shriver & Atkins' Inorganic Chemistry 5th Ed. Oxford University Press (2010). 5. Cotton, F.A., Wilkinson, G. and Gaus, P.L., Basic Inorganic Chemistry 3rd Ed.; Wiley India. 6. Sharpe, A.G., Inorganic Chemistry, 4th Indian Reprint (Pearson Education) 2005. 7. Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006. 8. Atkins, P.W. & Paula, J. Physical Chemistry, Oxford Press, 2006. 9. Mingos, D.M.P., Essential trends in inorganic chemistry. Oxford University Press (1998). 10. Winter, M. J., The Orbitron, http winter.group.shef.ac.uk/orbitron/ (2002). An illustrated gallery of atomic and molecular orbitals. 11. Burgess, J., Ions in solution: basic principles of chemical interactions. Ellis Horwood (1999)

		Credit
Practical		
Acid and Base Titrations 1. Estimation of carbonate and hydroxide present together in mixture 2. Estimation of carbonate and bicarbonate present together in a mixture. 3. Estimation of free alkali present in different soaps/detergents. **Oxidation-Reduction Titrimetric** 1. Estimation of Fe(II) using standardized $KMnO_4$ solution 2. Estimation of oxalic acid and sodium oxalate in a given mixture 3. Estimation of Fe(II) and Fe(III) in a given mixture using $K_2Cr_2O_7$ solution. 4. Estimation of Fe(III) and Mn(II) in a mixture using standardized $KMnO_4$ solution 5. Estimation of Fe(III) and Cu(II) in a mixture using $K_2Cr_2O_7$. 6. Estimation of Fe(III) and Cr(III) in a mixture using $K_2Cr_2O_7$. **Reference Books:** Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009		
CEMHT-4	**Organic Chemistry II**	**4 Credit**
Theory: Stereochemistry, General Treatment of Reaction Mechanism, Substitution and Elimination Reactions		
1. Stereochemistry-II: (15L) **Chirality arising out of Stereo-axis:** Stereoisomerism of substituted cumulenes with even and odd number of double bonds; chiral axis in allenes, spiro compounds, alkylidenecycloalkanes and biphenyls; related configurational descriptors (R_a/S_a and P/M); atropisomerism; racemisation of chiral biphenyls; buttressing effect. **Concept of prostereoisomerism:** Prostereogenic centre; concept of pro-chirality: topicity of ligands and faces (elementary idea); pro-R/pro-S, pro-E/pro-Z and Re/Si descriptors; pro-r and pro- s descriptors of ligands on pro-pseudoasymmetric centre. Conformation: **Conformational nomenclature:** eclipsed, staggered, gauche, syn and anti; dihedral angle, torsion angle; Klyne-Prelog terminology; P/M descriptors; energy barrier of rotation, concept of torsional and steric strains; relative stability of conformers on the basis of steric effect, dipole-dipole interaction and H-bonding; butane gauche interaction; conformational analysis of ethane, propane, n-butane, 2- methylbutane and 2,3- dimethylbutane; haloalkane, 1,2-dihaloalkanes and 1,2- diols (up to four carbons); 1,2-halohydrin; conformation of conjugated systems (s-cis and s-trans).		

2. General Treatment of Reaction Mechanism II: (18L)

Reaction thermodynamics:

Free energy and equilibrium, enthalpy and entropy factor, calculation of enthalpy change via BDE, intermolecular & intramolecular reactions.

Concept of organic acids and bases:

Effect of structure, substituent and solvent on acidity and basicity; proton sponge; gas-phase acidity and basicity; comparison between nucleophlicity and basicity; HSAB principle; application of thermodynamic principles in acid-base equilibria.

Tautomerism: Prototropy (keto-enol, nitro - aci-nitro, nitroso-oximino, diazo-amino and enamine-imine systems); valence tautomerism and ring-chain tautomerism; composition of the equilibrium in different systems (simple carbonyl; 1,2- and 1,3-dicarbonyl systems, phenols and related systems), factors affecting keto-enol tautomerism; application of thermodynamic principles in tautomeric equilibria.

Reaction kinetics: Rate constant and free energy of activation; concept of order and molecularity; free energy profiles for one-step, two-step and three-step reactions; catalyzed reactions: electrophilic and nucleophilic catalysis; kinetic control and thermodynamic control of reactions; isotope effect: primary and secondary kinetic isotopic effect (k_H/k_D); principle of microscopic reversibility; Hammond's postulate.

3. Substitution and Elimination Reactions: (27L)

Free-radical substitution reaction:

Halogentaion of alkanes, mechanism (with evidence) and stereochemical features; reactivity-selectivity principle in the light of Hammond's postulate.

Nucleophilic substitution reactions:

Substitution at sp3 centre: mechanisms (with evidence), relative rates & stereochemical features: S_N1, S_N2, S_N2', S_N1' (allylic rearrangement) and S_Ni; effects of solvent, substrate structure, leaving group and nucleophiles (including ambident nucleophiles, cyanide & nitrite); substitutions involving NGP; role of crown ethers and phase transfer catalysts; [systems: alkyl halides, allyl halides, benzyl halides, alcohols, ethers, epoxides]. Concept of aliphatic electrophilic substitution reactions (S_E1, S_E2, S_Ei).

Elimination reactions:

E1, E2, E1cb and Ei (pyrolytic syn eliminations); formation of alkenes and alkynes; mechanisms (with evidence), reactivity, regioselectivity (Saytzeff/Hofmann) and stereoselectivity; comparison between substitution and elimination; importance of Bredt's rule relating to the formation of C=C.

Reference Books:

1.Clayden, J., Greeves, N., Warren, S. Organic Chemistry, Second edition, Oxford University Press 2012. 2. Sykes, P. A guidebook to Mechanism in Organic Chemistry, Pearson Education, 2003. 3. Smith, J. G. Organic Chemistry, Tata McGraw-Hill Publishing Company Limited. 4. Carey, F. A. & Guiliano, R. M. Organic Chemistry, Eighth edition, McGraw Hill Education, 2012. 5. Loudon, G.M. Organic Chemistry, Fourth edition, Oxford University Press, 2008. 6. Eliel, E.L. & Wilen, S. H. Stereochemistry of Organic Compounds,

Wiley: London, 1994. 7. Nasipuri, D. Stereochemistry of Organic Compounds, Wiley Eastern Limited. 8. Morrison, R. N. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 9. Finar, I. L. Organic Chemistry (Volume 1) Pearson Education. 10. Graham Solomons, T.W., Fryhle, C. B. Organic Chemistry, John Wiley & Sons, Inc. 11. Eames, J., Peach, J. M. Stereochemistry at a Glance, Blackwell Publishing, 2003. 12. Robinson, M. J. T. Stereochemistry, Oxford Chemistry Primer, Oxford University Press, 2005. 13. Maskill, H. Mechanisms of Organic Reactions, Oxford Chemistry Primer, Oxford University Press. 14. March, J. Advanced Organic Chemistry, Fourth edition, Wiley.

CEMHP-4	**Organic Chemistry – II**	**2 Credit**

Practical

Organic Preparations:

A. The following reactions are to be performed, noting the yield of the crude product:

1. Nitration of aromatic compounds
2. Condensation reactions
3. Hydrolysis of amides/imides/esters
4. Acetylation of phenols/aromatic amines
5. Benzoylation of phenols/aromatic amines
6. Side chain oxidation of aromatic compounds
7. Diazo coupling reactions of aromatic amines
8. Bromination of anilides using green approach (Bromate-Bromide method)
9. Redox reaction including solid-phase method
10. Green ‘multi-component-coupling’ reaction
11. Selective reduction of m-dinitrobenzene to m-nitroaniline

Students must also calculate percentage yield, based upon isolated yield (crude) and theoretical yield.

B. Purification of the crude product is to be made by crystallization from water/alcohol, crystallization after charcoal treatment, or sublimation, whichever is applicable.

C. Melting point of the purified product is to be noted.

Reference Books:

1.Vogel, A. I. Elementary Practical Organic Chemistry, Part 1: Small scale Preparations, CBS Publishers and Distributors. 2. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N. University of Calcutta, 2003. 3. Mann, F.G. & Saunders, B.C. Practical Organic Chemistry, Pearson Education (2009).4. Furniss, B.S., Hannaford, A.J., Smith, P.W.G. & Tatchell, A.R. Practical Organic Chemistry, 5th Ed. Pearson (2012). 5. Ahluwalia, V.K. & Aggarwal, R. Comprehensive Practical Organic Chemistry: Preparation and Quantitative Analysis, University Press (2000). 6. Practical Workbook Chemistry (Honours), UGBS, Chemistry, University of Calcutta, 2015.

Semester - III		
CEMHT-5	**Physical Chemistry – II**	**4 Credit**
Theory: Transport Processes, Applications of Thermodynamics – I, Foundation of Quantum Mechanics		

1. Transport Processes: (20L)

Fick's law: Flux, force, phenomenological coefficients & their inter-relationship (general form), different examples of transport properties

Viscosity: General features of fluid flow (streamline flow and turbulent flow); Newton's equation, viscosity coefficient; Poiseuille's equation; Principle of determination of viscosity coefficient of liquids by falling sphere method; Temperature variation of viscosity of liquids and comparison with that of gases.

Conductance and transport number: Ion conductance; Conductance and measurement of conductance, cell constant, specific conductance and molar conductance; Variation of specific and equivalent conductance with dilution for strong and weak electrolytes; Kohlrausch's law of independent migration of ions; Equivalent and molar conductance at infinite dilution and their determination for strong and weak electrolytes; Debye–Huckel theory of Ion atmosphere (qualitative)-asymmetric effect, relaxation effect and electrophoretic effect; Ostwald's dilution law; Ionic mobility; Application of conductance measurement (determination of solubility product and ionic product of water); Conductometric titrations. Transport number, Principles of Hittorf's and Moving-boundary method; Wien effect, Debye-Falkenhagen effect, Walden's rule

2. Applications of Thermodynamics –I (20L)

Partial properties and Chemical potential: Chemical potential and activity, partial molar quantities, relation between Chemical potential and Gibb's free energy and other thermodynamic state functions; variation of Chemical potential (μ) with temperature and pressure; Gibbs-Duhem equation; fugacity and fugacity coefficient; Variation of thermodynamic functions for systems with variable composition; Equations of states for these systems, Change in G, S H and V during mixing for binary solutions.

Chemical Equilibrium: Thermodynamic conditions for equilibrium, degree of advancement; van't Hoff's reaction isotherm (deduction from chemical potential); Variation of free energy with degree of advancement; Equilibrium constant and standard Gibbs free energy change; Definitions of KP, KC and KX; van't Hoff's reaction isobar and isochore from different standard states; Shifting of equilibrium due to change in external parameters e.g. temperature and pressure; variation of equilibrium constant with addition to inert gas; Le Chatelier's principle and its derivation.

Nernst's distribution law: Application- (finding out Keq using Nernst distribution law for $KI+I_2 = KI_3$ and dimerization of benzene.

Chemical potential and other properties of ideal substances- pure and mixtures:

Pure ideal gas-its Chemical potential and other thermodynamic functions and their changes during a change of; Thermodynamic parameters of mixing;

Chemical potential of an ideal gas in an ideal gas mixture; Concept of standard states and choice of standard states of ideal gases

Condensed_Phase – Chemical potential of pure solid and pure liquids, Ideal solution – Definition, Raoult's law; Mixing properties of ideal solutions, chemical potential of a component in an ideal solution; Choice of standard states of solids and liquids

3. **Foundation of Quantum Mechanics (20L)**

Beginning of Quantum Mechanics: Wave-particle duality, light as particles: photoelectric and Compton effects; electrons as waves and the de Broglie hypothesis; Uncertainty relations (without proof)

Wave function: Schrodinger time-independent equation; nature of the equation, acceptability conditions imposed on the wave functions and probability interpretations of wave function

Concept of Operators: Elementary concepts of operators, eigenfunctions and eigenvalues; Linear operators; Commutation of operators, commutator and uncertainty relation; Expectation value; Hermitian operator; Postulates of Quantum Mechanics

Particle in a box: Setting up of Schrodinger equation for one- dimensional box and its solution; Comparison with free particle eigenfunctions and eigenvalues. Properties of PB wave functions (normalisation, orthogonality, probability distribution); Expectation values of x, x2, px and px2 and their significance in relation to the uncertainty principle; Extension of the problem to two and three dimensions and the concept of degenerate energy levels

Simple Harmonic Oscillator: setting up of the Schrodinger stationary equation, energy expression (without derivation), expression of wave function for n = 0 and n = 1 (without derivation) and their characteristic features.

Reference Books

1. Atkins, P. W. & Paula, J. de Atkins', Physical Chemistry, Oxford University Press. Castellan, G. W. Physical Chemistry, Narosa. 3. McQuarrie, D. A. & Simons, J. D. Physical Chemistry: A Molecular Approach, Viva Press. 4. Levine, I. N. Physical Chemistry, Tata McGraw-Hill. 4. Rakshit, P.C., Physical Chemistry, Sarat Book House. 5. Moore, W. J. Physical Chemistry, Orient Longman. 6. Mortimer, R. G. Physical Chemistry, Elsevier. 7. Denbigh, K. The Principles of Chemical Equilibrium Cambridge University Press. 8. Engel, T. & Reid, P. Physical Chemistry, Pearson. 9. Levine, I. N. Quantum Chemistry, PHI. 10. Atkins, P. W. Molecular Quantum Mechanics, Oxford. 11. emansky, M. W. & Dittman, R.H. Heat and Thermodynamics, Tata-McGraw-Hill. 12. Rastogi, R. P. & Misra, R.R. An Introduction to Chemical Thermodynamics, Vikas. 13. Klotz, I.M., Rosenberg, R. M. Chemical Thermodynamics:Basic Concepts and Methods Wiley. 14. Glasstone, S. An Introduction to Electrochemistry, East-West Press.

CEMHP-5	Physical Chemistry – II	2 Credit

Practical

1.Conductometric titration of an acid (strong, weak/ monobasic, dibasic) against strong base.

2.Study of saponification reaction conductometrically.

3.Verification of Ostwald's dilution law and determination of K_a of weak acid.

4.Study of viscosity of unknown liquid (glycerol, sugar) with respect to water.

5.Determination of partition coefficient for the distribution of I_2 between water and CCl_4.

6.Determination of K_{eq} for $KI + I_2 = KI_3$, using partition coefficient between water and CCl_4

Reference Books

1. Viswanathan, B., Raghavan, P.S. Practical Physical Chemistry Viva Books (2009) 2. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson. 3. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman (2007). 4. Palit, S.R., De, S. K. Practical Physical Chemistry Science Book Agency. 5. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N., University of Calcutta. 6 .Levitt, B. P. edited Findlay's Practical Physical Chemistry Longman Group Ltd. 7. Gurtu, J. N., Kapoor, R., Advanced Experimental Chemistry S. Chand & Co..Ltd.

CEMHT-6	Inorganic Chemistry – II	4 Credit

Theory: Chemical Bonding-I, Chemical Bonding-II, Radioactivity

1. Chemical Bonding-I (21L)

Ionic bond: General characteristics, types of ions, size effects, radius ratio rule and its application and limitations. Packing of ions in crystals. Born-Landé equation with derivation and importance of Kapustinskii expression for lattice energy. Madelung constant, Born-Haber cycle and its application, Solvation energy. Defects in solids (elementary idea). Solubility energetics of dissolution process.

Covalent bond: Polarizing power and polarizability, ionic potential, Fazan's rules. Lewis structures, formal charge. Valence Bond Theory. The hydrogen molecule (Heitler-London approach), directional character of covalent bonds, hybridizations, equivalent and nonequivalent hybrid orbitals, Bent's rule, Dipole moments, VSEPR theory, shapes of molecules and ions containing lone pairs and bond pairs (examples from main groups chemistry) and multiple bonding (σ and π bond approach).

2. Chemical Bonding-II (21L)

Molecular orbital concept of bonding (The approximations of the theory, Linear combination of atomic orbitals (LCAO)) (elementary pictorial approach): sigma and pibonds and delta interaction, multiple bonding. Orbital designations: gerade, ungerade, HOMO, LUMO. Orbital mixing, MO diagrams of H_2, Li_2, Be_2, B_2, C_2, N_2, O_2, F_2, and their ions wherever possible; Heteronuclear molecular orbitals: CO, NO, NO^+, CN^-, HF, BeH_2, CO_2 and H_2O. Bond properties: bond orders, bond lengths.

Metallic Bond: Qualitative idea of valence bond and band theories. Semiconductors and insulators, defects in solids.

Weak Chemical Forces: van der Waals forces, ion-dipole forces, dipole-dipole interactions, induced dipole interactions, Instantaneous dipole-induced dipole interactions. Repulsive forces, Intermolecular forces: Hydrogen bonding (theories of hydrogen bonding, valence bond treatment), receptor-guest interactions, Halogen bonds. Effects of chemical force, melting and boiling points.

3. Radioactivity (18L)

Nuclear stability and nuclear binding energy. Nuclear forces: meson exchange theory. Nuclear models (elementary idea): Concept of nuclear quantum number, magic numbers.

Nuclear Reactions: Artificial radioactivity, transmutation of elements, fission, fusion and spallation. Nuclear energy and power generation. Separation and uses of isotopes.

Radio chemical methods: principles of determination of age of rocks and minerals, radio carbon dating, hazards of radiation and safety measures

Reference Books

1. Lee, J. D. Concise Inorganic Chemistry 5th Ed., John Wiley and sons 2008. 2. Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006. 3. Douglas, B.E. and McDaniel, D.H. Concepts & Models of Inorganic Chemistry Oxford, 1970. 4. Porterfield, H. W., Inorganic Chemistry, Second Edition, Academic Press, 2005. 5. Purecell, K.F. and Kotz, J.C., An Introduction to Inorganic Chemistry, Saunders: Philadelphia, 1980. 6. Cotton, F.A., Wilkinson, G., & Gaus, P.L. Basic Inorganic Chemistry 3rd Ed.; Wiley India. 7. Gillespie, R. J. and Hargittai, I., The VSEPR Model of Molecular Geometry, Prentice Hall (1992). 8. Albright, T., Orbital interactions in chemistry, John Wiley and Sons (2005). 9. Mingos, D.M.P., Essential trends in inorganic chemistry. Oxford University Press (1998). 10. Miessler, G. L., Fischer, P. J., Tarr, D. A., Inorganic Chemistry, Pearson, 5th Edition.

CEMHP-6	**Inorganic Chemistry – II**	**2 Credit**
Practical		
Iodo / Iodimetric Titrations 1. Estimation of Cu(II) 2. Estimation of Vitamin C 3. Estimation of (i) arsenite and (ii) antimony in tartar-emetic iodimetrically 4. Estimation of available chlorine in bleaching powder **Estimation of metal content in some selective samples** 1. Estimation of Cu in brass. 2. Estimation of Cr and Mn in Steel. 3. Estimation of Fe in cement. **Reference Books** Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009.		

CEMHT-7	**Organic Chemistry – III**	**4 Credit**
Theory: Chemistry of alkenes and alkynes, Aromatic Substitution, Carbonyl and Related Compounds, Organometallics		
1. Chemistry of alkenes and alkynes: (15L) **Addition to C=C:** mechanism (with evidence wherever applicable), reactivity, regioselectivity (Markownikoff and anti- Markownikoff additions) and stereoselectivity; reactions: hydrogenation, halogenations, iodolactonisation, hydrohalogenation, hydration, oxymercuration-demercuration, hydroboration-oxidation, epoxidation, syn and anti-hydroxylation, ozonolysis, addition of singlet and triplet carbenes; electrophilic addition to diene (conjugated dienes and allene); radical addition: HBr addition; mechanism of allylic and benzylic bromination in competition with brominations across C=C; use of NBS; dissolving metal reduction of alkenes; interconversion of E- and Z- alkenes; contra-thermodynamic isomerization of internal alkenes. **Addition to C≡C (in comparison to C=C):** mechanism, reactivity, regioselectivity (Markownikoff and anti-Markownikoff addition) and stereoselectivity; reactions: hydrogenation, halogenations, hydrohalogenation, hydration, oxymercuration-demercuration, hydroboration-oxidation, dissolving metal reduction of alkynes (Birch); reactions of terminal alkynes by exploring its acidity; interconversion of terminal and non-terminal alkynes. **2. Aromatic Substitutions: (9L)** **Electrophilic aromatic substitution:** mechanisms and evidences in favour of it; orientation and reactivity; reactions: nitration, nitrosation, sulfonation, halogenation, Friedel-Crafts reaction; one-carbon electrophiles (reactions: chloromethylation, Gatterman-Koch, Gatterman,		

Houben- Hoesch, Vilsmeier-Haack, Reimer-Tiemann, Kolbe-Schmidt); Ipso substitituion.

Nucleophilic aromatic substitution: addition-elimination mechanism and evidences in favour of it; S_NAr mechanism; cine substitution (benzyne mechanism), structure of benzyne.

3. Carbonyl and Related Compounds: (30L)

Addition to C=O: structure, reactivity and preparation of carbonyl compounds; mechanism (with evidence), reactivity, equilibrium and kinetic control; Burgi-Dunitz trajectory in nucleophilic additions; formation of hydrates, cyano hydrins and bisulphite adduct;nucleophilic addition-elimination reactions with alcohols, thiols and nitrogen- based nucleophiles; reactions:benzoin condensation, Cannizzaro and Tischenko reactions, reactions with ylides: Wittig and Corey-Chaykovsky reaction; Rupe rearrangement, oxidations and reductions: Clemmensen, Wolff-Kishner, $LiAlH_4$, $NaBH_4$, MPV, Oppenauer, Bouveault-Blanc, acyloin condensation; oxidation of alcohols with PDC and PCC; periodic acid and lead tetraacetate oxidation of 1,2-diols.

Exploitation of acidity of α-H of C=O: formation of enols and enolates; kinetic and thermodynamic enolates; reactions (mechanism with evidence): halogenation of carbonyl compounds under acidic and basic conditions, Hell-Volhard-Zelinsky (H. V. Z.) reaction, nitrosation, SeO_2 (Riley) oxidation; condensations (mechanism with evidence): Aldol, Knoevenagel, Claisen-Schmidt, Claisen ester including Dieckmann, Stobbe; Mannich reaction, Perkin reaction, Favorskii rearrangement; alkylation of active methylene compounds; preparation and synthetic applications of diethyl malonate and ethyl acetoacetate; specific enol equivalents (lithium enolates, enamines, aza-enolates and silyl enol ethers) in connection with alkylation, acylation and aldol type reaction.

Elementary ideas of Green Chemistry: Twelve (12) principles of green chemistry; planning of green synthesis; common organic reactions and their counterparts: reactions: Aldol, Friedel-Crafts, Michael, Knoevenagel, Cannizzaro, benzoin condensation and Dieckmann condensation.

Nucleophilic addition to α,β-unsaturated carbonyl system: general principle and mechanism (with evidence); direct and conjugate addition, addition of enolates (Michael reaction), Stetter reaction, Robinson annulations.

Substitution at sp^2 carbon (C=O system): mechanism (with evidence): $B_{AC}2$, $A_{AC}2$, $A_{AC}1$, $A_{AL}1$ (in connection to acid and ester); acid derivatives: amides, anhydrides & acyl halides (formation and hydrolysis including comparison).

4. Organometallics: (6L)

Grignard reagent; Organolithiums; Gilman cuprates: preparation and

reactions (mechanism with evidence); addition of Grignard and organolithium to carbonyl compounds; substitution on -COX; directed ortho metalation of arenes using organolithiums, conjugate addition by Gilman cuprates; Corey-House synthesis; abnormal behavior of Grignard reagents; comparison of reactivity among Grignard, organolithiums and organocopper reagents; Reformatsky reaction; Blaise reaction; concept of umpolung and base-nucleophile dichotomy in case of organometallic reagents.

Reference Books:

1.Clayden, J., Greeves, N., Warren, S. Organic Chemistry, Second edition, Oxford University Press 2012. 2. Sykes, P. A guidebook to Mechanism in Organic Chemistry, Pearson Education, 2003. 3. Smith, J. G. Organic Chemistry, Tata McGraw-Hill Publishing Company Limited. 4. Carey, F. A., Guiliano, R. M. Organic Chemistry, Eighth edition, McGraw Hill Education, 2012. 5. Loudon, G. M. Organic Chemistry, Fourth edition, Oxford University Press, 2008. 6. Norman, R.O. C., Coxon, J. M. Principles of Organic Synthesis, Third Edition, Nelson Thornes, 2003. 7. Morrison, R. N. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 8. Finar,I. L. Organic Chemistry (Volume 1), Pearson Education. 9. Graham Solomons, T.W., Fryhle, C. B. Organic Chemistry, John Wiley & Sons, Inc. 10. March, J. Advanced Organic Chemistry, Fourth edition, Wiley. 11. Jenkins, P. R., Organometallic Reagents in Synthesis, Oxford Chemistry Primer, Oxford University Press. 12. Ward, R. S., Bifunctional Compounds, Oxford Chemistry Primer, Oxford University Press

CEMHP-7	**Organic Chemistry – III**	**2 Credit**

Practical

Qualitative Analysis of Single Solid Organic Compounds:

1. Detection of special elements (N, S, Cl, Br) by Lassaigne's test
2. Solubility and classification (solvents: H_2O, 5% HCl, 5% NaOH and 5% $NaHCO_3$)
3. Detection of the following functional groups by systematic chemical tests:
4. Aromatic amino ($Ar\text{-}NH_2$), aromatic nitro ($Ar\text{-}NO_2$), amido ($\text{-}CONH_2$, including imide), phenolic hydroxyl (Ph–OH), carboxylic acid (-COOH), carbonyl (-CHO and >C=O); only one test for each functional group is to be reported.
5. Melting point of the given compound
6. Preparation, purification and melting point determination of a crystalline derivative of the given compound
7. Identification of the compound through literature survey.

Each student, during laboratory session, is required to carry out qualitative chemical tests for all the special elements and the functional groups with relevant derivatization in known and unknown (at least six) organic compounds

Reference Books:

Vogel, A. I. Elementary Practical Organic Chemistry, Part 2: Qualitative Organic Analysis, CBS Publishers and Distributors. 2. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N. University of Calcutta, 2003. 3. Mann, F.G. & Saunders, B.C. Practical Organic Chemistry, Pearson Education (2009). 4. Furniss, B.S., Hannaford, A.J., Smith, P.W.G., Tatchell, A.R. Practical Organic Chemistry, 5th Ed., Pearson (2012). 5. Clarke,H. T., A Handbook of Organic Analysis (Qualitative and Quantitative), Fourth Edition, CBS Publishers and Distributors (2007). 6. Practical Workbook Chemistry (Honours), UGBS, Chemistry, University of Calcutta, 2015.

Semester - IV

CEMHT-8	**Physical-III**	**4 Credit**

Application of Thermodynamics- II, Electrical properties of molecules, Quantum Chemistry

1. Application of Thermodynamics – II (20 L)

Colligative properties: Vapour pressure of solution; Ideal solutions, ideally dilute solutions and colligative properties; Raoult's law; Thermodynamic derivation using chemical potential to derive relations between the four colligative properties [(i) relative lowering of vapour pressure, (ii) elevation of boiling point, (iii) Depression of freezing point,(iv) Osmotic pressure and amount of solute. Applications in calculating molar masses of normal, dissociated and associated solutes in solution; Abnormal colligative properties.

Phase rule: Definitions of phase, component and degrees of freedom; Phase rule and its derivations; Definition of phase diagram; Phase diagram for water, CO_2, Sulphur. First order phase transition and Clapeyron equation; Clausius-Clapeyron equation derivation and use; Liquid vapour equilibrium for two component systems; Phenol- water system. Three component systems, water-chloroform-acetic acid system, triangular plots.
Binary solutions: Ideal solution at fixed temperature and pressure; Principle of fractional distillation; Duhem-Margules equation; Henry's law; Konowaloff's rule; Positive and negative deviations from ideal behavior; Azeotropic solution; Liquid- liquid phase diagram using phenol-water system; Solid-liquid phase diagram; Eutectic mixture.

2. Electrical Properties of molecules (20 L)

Ionic equilibria: Chemical potential of an ion in solution; Activity and activity coefficients of ions in solution; Debye-Huckel limiting law-brief qualitative description of the postulates involved, qualitative idea of the model, the equation (without derivation) for ion-ion atmosphere

interaction potential. Estimation of activity coefficient for electrolytes using Debye-Huckel limiting law; Derivation of mean ionic activity coefficient from the expression of ion-atmosphere interaction potential; Applications of the equation and its limitations.

Electromotive Force: Quantitative aspects of Faraday's laws of electrolysis, rules of oxidation/reduction of ions based on half-cell potentials, applications of electrolysis in metallurgy and industry; Chemical cells, reversible and irreversible cells with examples; Electromotive force of a cell and its measurement, Nernst equation; Standard electrode (reduction) potential and its application to different kinds of half- cells. Application of EMFmeasurements in determining (i) free energy, enthalpy and entropy of a cell reaction, (ii) equilibrium constants, and (iii) pH values, using hydrogen, quinone-hydroquinone, glass electrodes.
Concentration cells with and without transference, liquid junction potential; Determination of activity coefficients and transference numbers; Qualitative discussion of potentiometric titrations (acid-base, redox, precipitation).

Dipole moment and polarizability: Polarizability of atoms and molecules, dielectric constant and polarisation, molar polarisation for polar and non-polar molecules; Clausius-Mosotti equation and Debye equation (both without derivation) and their application; Determination of dipole moments.

3. Quantum Chemistry (20 L)

Angular momentum: Commutation rules, quantization of square of total angular momentum and z-component; Rigid rotator model of rotation of diatomic molecule; Schrödinger equation, transformation to spherical polar coordinates; Separation of variables. Spherical Harmonics, Discussion of solution.

Qualitative treatment of hydrogen atom and hydrogen-like ions: Setting up of Schrödinger equation in spherical polar coordinates, radial part, quantization of energy (only final energy expression); Average and most probable distances of electron from nucleus; Setting up of Schrödinger equation for many-electron atoms (He, Li).

LCAO and HF-SCF: Covalent bonding, valence bond and molecular orbital approaches, LCAO-MO treatment of H_2^+; Bonding and antibonding orbitals; Qualitative extension to H_2; Comparison of LCAO-MO and VB treatments of H_2 and their limitations; Hartree-Fock method development, SCF and configuration interaction (**only basics**).

Reference Books

1.Castellan, G. W. Physical Chemistry, Narosa. 2. Atkins, P. W. & Paula, J. de Atkins', Physical Chemistry, Oxford University Press. 3. McQuarrie, D. A. & Simons, J. D. Physical Chemistry: A Molecular Approach, VivaPress. 4. Levine, I. N. Physical Chemistry, Tata McGraw-Hill . 5. Moore, W. J. Physical Chemistry, Orient Longman. 6. Mortimer, R. G. Physical Chemistry, Elsevier. 7. Engel, T. & Reid, P. Physical Chemistry, Pearson. 8. Levine, I. N. Quantum Chemistry, PHI. 9. Atkins, P. W. Molecular Quantum Mechanics, Oxford. 10. Engel, T. & Reid, P. Physical Chemistry, Pearson. 11. Maron, S.H., Prutton, C. F., Principles of Physical Chemistry, McMillan. 12. Klotz, I.M., Rosenberg, R. M. Chemical Thermodynamics:Basic Concepts and MethodsWiley. 13. Rastogi, R. P. &Misra, R.R. An Introduction to Chemical Thermodynamics, Vikas. 14

CEMHP-8	**Physical-III**	**2 Credit**

Practical

1. Determination of solubility of sparingly soluble salt in water, in electrolyte with common ions and in neutral electrolyte (using common indicator)
2. Potentiometric titration of Mohr's salt solution against standard $K_2Cr_2O_7$ solution
3. Determination of K_{sp} for AgCl by potentiometric titration of $AgNO_3$ solution against standard KCl solution
4. Effect of ionic strength on the rate of Persulphate – Iodide reaction
5. Study of phenol-water phase diagram
6. pH-metric titration of acid (mono- and di-basic) against strong base

Reference Books

Viswanathan, B., Raghavan, P.S. Practical Physical Chemistry Viva Books (2009) 2. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson. 3. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman (2007). 4. Palit, S.R., De, S. K. Practical Physical Chemistry Science Book Agency. 5. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N., University of Calcutta. 6 .Levitt, B. P. edited Findlay's Practical Physical Chemistry Longman Group Ltd. 7. Gurtu, J. N., Kapoor, R., Advanced Experimental Chemistry
S. Chand & Co..Ltd

CEMHT-9	**Inorganic Chemistry – III**	**4 Credit**

Theory: General Principles of Metallurgy, Chemistry of s and p Block Elements, Noble Gases, Inorganic Polymers, Coordination Chemistry-I

1. **General Principles of Metallurgy: (15L)**

Chief modes of occurrence of metals based on standard electrode potentials. Ellingham diagrams for reduction of metal oxides using carbon and carbon monoxide as reducing agent. Electrolytic Reduction,

Hydrometallurgy. Methods of purification of metals: Electrolytic Kroll process, Parting process, van Arkel-de Boer process and Mond's process, Zone refining.

2. Chemistry of s and p Block Elements: (10L)

Relative stability of different oxidation states, diagonal relationship and anomalous behaviour of first member of each group. Allotropy and catenation. Study of the following compounds with emphasis on structure, bonding, preparation, properties and uses. Beryllium hydrides and halides. Boric acid and borates, boron nitrides, borohydrides (diborane) and graphitic compounds, silanes, Oxides and oxoacids of nitrogen, phosphorus, sulphur and chlorine. Peroxo acids of sulphur, sulphur- nitrogen compounds, interhalogen compounds, polyhalide ions, pseudohalogens, Fluorocarbons and basic properties of halogens.

3. Noble Gases: (10L)

Occurrence and uses, rationalization of inertness of noble gases, Clathrates; preparation and properties of XeF_2, XeF_4 and XeF_6; Nature of bonding in noble gas compounds (Valence bond treatment and MO treatment for XeF_2 and XeF_4). Xenon-oxygen compounds. Molecular shapes of noble gas compounds (VSEPR theory).

4. Inorganic Polymers: (10L)

Types of inorganic polymers, comparison with organic polymers, synthesis, structural aspects and applications of silicones and siloxanes. Borazines, silicates and phosphazenes.

5. Coordination Chemistry-I: (15L)

Coordinate bonding: double and complex salts. Werner's theory of coordination complexes, Classification of ligands, Ambidentate ligands, chelates, Coordination numbers, IUPAC nomenclature of coordination complexes (up to two metal centers), Isomerism in coordination compounds, constitutional and stereo isomerism, Geometrical and optical isomerism in square planar and octahedral complexes.

Reference Books

1.Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006. 2. Greenwood,N.N. & Earnshaw A. Chemistry of the Elements, Butterworth-Heinemann, 1997. 3. Cotton, F.A., Wilkinson, G., Murrillo, C. A., Bochmann, M., Advanced Inorganic Chemistry 6th Ed. 1999., Wiley. 4. Miessler, G. L. & Donald, A. Tarr. Inorganic Chemistry 4th Ed., Pearson, 2010. 5. Purecell, K.F. and Kotz, J.C., An Introduction to Inorganic Chemistry, Saunders: Philadelphia, 1980. 6. Mingos, D.M.P., Essential trends in inorganic chemistry. Oxford University Press (1998). 7. Anil J. Elias, The Chemistry of *p*-Block elements: Synthese, Reactions and Applications.

CEMHP-9	Inorganic Chemistry – III	2 Credit
Practical		

Complexometric titration

1. Zn(II)
2. Zn(II) in a Zn(II) and Cu(II) mixture.
3. Ca(II) and Mg(II) in a mixture.
4. Hardness of water.

Inorganic preparations

1. $[Cu(CH_3CN)_4]PF_6/ClO_4$
2. Cis and trans $K[Cr(C_2O_4)_2(H_2O)_2]$
3. $[Cu(CH_3CN)_4]PF_6/ClO_4$
4. Cis and trans $K[Cr(C_2O_4)_2(H_2O)_2]$
5. Potassium dioxalatodiaquachromate(III)
6. Tetraamminecarbonatocobalt (III) ion
7. Potassium tris(oxalate)ferrate(III)
8. Tris-(ethylenediamine) nickel(II) chloride.
9. $[Mn(acac)_3]$ and $Fe(acac)_3]$ (acac= acetylacetonate)

Reference Books

Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009.

CEMHT-10	Organic Chemistry – IV	4 Credit

Theory: Nitrogen compounds, Rearrangements, The Logic of Organic Synthesis, Organic Spectroscopy.

1. **Nitrogen compounds: (9L)**

   Amines: Aliphatic & Aromatic: preparation, separation (Hinsberg's method) and identification of primary, secondary and tertiary amines; reaction (with mechanism): Eschweiler–Clarke methylation, diazo coupling reaction, Mannich reaction; formation and reactions of phenylenediamines, diazomethane and diazoacetic ester.

   Nitro compounds (aliphatic and aromatic): preparation and reaction (with mechanism): reduction under different conditions; Nef carbonyl synthesis, Henry reaction and conjugate addition of nitroalkane anion.

   Alkylnitrile and isonitrile: preparation and reaction (with mechanism): Thorpe nitrile condensation, von Richter reaction.

   Diazonium salts and their related compounds: reactions (with mechanism) involving replacement of diazo group; reactions: Gomberg, Meerwein, Japp-Klingermann.

2. **Rearrangements: (11L)**

   Mechanism with Evidence and Stereochemical Features for the Following Rearrangement to electron-deficient carbon: Wagner-

Meerwein rearrangement, pinacol rearrangement, dienone-phenol; Wolff rearrangement in Arndt-Eistert synthesis, benzil-benzilic acid rearrangement, Demjanov rearrangement, Tiffeneau–Demjanov rearrangement.

Rearrangement to electron-deficient nitrogen: rearrangements: Hofmann, Curtius, Lossen, Schmidt and Beckmann.

Rearrangement to electron-deficient oxygen: Baeyer-Villiger oxidation, cumene hydroperoxide-phenol rearrangement and Dakin reaction.

Aromatic rearrangements: Migration from oxygen to ring carbon: Fries rearrangement and Claisen rearrangement.

Migration from nitrogen to ring carbon: Hofmann-Martius rearrangement, Fischer-Hepp rearrangement, N-azo to C-azo rearrangement, Bamberger rearrangement, Orton rearrangement and benzidine rearrangement.

Rearrangement reactions by green approach: Fries rearrangement, Claisen rearrangement, Beckmann rearrangement, Baeyer-Villiger oxidation.

3. **The Logic of Organic Synthesis: (20L)**

Retrosynthetic analysis: disconnections; synthons, donor and acceptor synthons; natural reactivity and umpolung; latent polarity in bifunctional compounds: consonant and dissonant polarity; illogical electrophiles and nucleophiles; synthetic equivalents; functional group interconversion and addition (FGI and FGA); C-C disconnections and synthesis: one-group and two-group (1,2- to 1,5-dioxygenated compounds), reconnection (1,6-dicarbonyl); protection-deprotection strategy (alcohol, amine, carbonyl, acid).

Strategy of ring synthesis: thermodynamic and kinetic factors; synthesis of large rings, application of high dilution technique.

Asymmetric synthesis: stereoselective and stereospecific reactions; diastereoselectivity and enantioselectivity (only definition); enantioselectivity: kinetically controlled MPV reduction; diastereoselectivity: addition of nucleophiles to C=O adjacent to a stereogenic centre: Felkin-Anh and Zimmermann-Traxler models.

4. **Organic Spectroscopy: (20L)**

UV Spectroscopy: introduction; types of electronic transitions, end absorption; transition dipole moment and allowed/forbidden transitions; chromophores and auxochromes; Bathochromic and Hypsochromic shifts; intensity of absorptions (Hyper-/Hypochromic effects); application of Woodward's Rules for calculation of λmax for the following systems: conjugated diene, α,β-unsaturated aldehydes and ketones (alicyclic, homoannular and heteroannular); extended conjugated systems (dienes, aldehydes and ketones); relative positions of λmax considering conjugative effect, steric effect, solvent effect, effect of pH; effective chromophore concentration: keto-enol systems; benzenoid

transitions.

IR Spectroscopy: introduction; modes of molecular vibrations (fundamental and non-fundamental); IR active molecules; application of Hooke's law, force constant; fingerprint region and its significance; effect of deuteration; overtone bands; vibrational coupling in IR; characteristic and diagnostic stretching frequencies of C-H, N-H, O-H, C-O, C-N, C-X, C=C (including skeletal vibrations of aromatic compounds), C=O, C=N, N=O, C≡C, C≡N; characteristic/diagnostic bending vibrations are included; factors affecting stretching frequencies: effect of conjugation, electronic effects, mass effect, bond multiplicity, ring-size, solvent effect, H-bonding on IR absorptions; application in functional group analysis.

NMR Spectroscopy: introduction; nuclear spin; NMR active molecules; basic principles of Proton Magnetic Resonance; equivalent and non-equivalent protons; chemical shift and factors influencing it; ring current effect; significance of the terms: up-/downfield, shielded and deshielded protons; spin coupling and coupling constant (1st order spectra); relative intensities of first-order multiplets: Pascal's triangle; chemical and magnetic equivalence in NMR ; elementary idea about non-first-order splitting; anisotropic effects in alkene, alkyne, aldehydes and aromatics;
NMR peak area, integration; relativepeak positions with coupling patterns of common organic compounds (both aliphatic and benzenoid-aromatic); rapid proton exchange; interpretation of NMR spectra of simple compounds.
Applications of IR, UV and NMR spectroscopy for identification of simple organic molecules.

Reference Books

1.Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 2. Finar, I. L. Organic Chemistry (Volume 2: Stereochemistry and the Chemistry of Natural Products), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 3. Norman, R.O. C., Coxon, J. M. Principles of Organic Synthesis, Third Edition, Nelson Thornes, 2003. 4. Clayden, J., Greeves, N., Warren, S., Organic Chemistry, Second edition, Oxford University Press 2012. 5. Silverstein, R. M., Bassler, G. C., Morrill, T. C. Spectrometric Identification of Organic Compounds, John Wiley and Sons, INC, Fifth edition. 6. Kemp, W. Organic Spectroscopy, Palgrave. 6. Pavia, D. L. et al. Introduction to Spectroscopy, 5th Ed. Cengage Learning India Ed. (2015). 7. Dyer, J. Application of Absorption Spectroscopy of Organic Compounds, PHI Private Limited. 8. March, J. Advanced Organic Chemistry, Fourth edition, Wiley.9. Harwood, L. M., Polar Rearrangements, Oxford Chemistry Primer, Oxford University Press. 10. Bailey, Morgan, Organonitrogen Chemistry, Oxford Chemistry Primer, Oxford University Press. Warren, S. Organic Synthesis the Disconnection Approach, John Wiley and Sons. Warren, S., Designing Organic Synthesis, Wiley India, 2009. 13. Carruthers, W. Modern methods of Organic Synthesis, Cambridge University

Press. 14. Willis, C. A., Wills, M., Organic Synthesis, Oxford Chemistry Primer, Oxford University Press.

CEMHP-10	**Organic Chemistry – IV**	**2 Credit**

Practical

List of Practical

1. Estimation of glycine by Sörensen's formol method
2. Estimation of glucose by titration using Fehling's solution
3. Estimation of sucrose by titration using Fehling's solution
4. Estimation of vitamin-C (reduced)
5. Estimation of aromatic amine (aniline) by bromination (Bromate-Bromide) method
6. Estimation of phenol by bromination (Bromate-Bromide) method
7. Estimation of formaldehyde (Formalin)
8. Estimation of acetic acid in commercial vinegar
9. Estimation of urea (hypobromite method)
10. Estimation of saponification value of oil/fat/ester

Reference Books:

1.Arthur, I. V. Quantitative Organic Analysis, Pearson, 2. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N., University of Calcutta

Semester - V

CEMHT-11	**Inorganic- IV**	**4 Credit**

Theory: Coordination Chemistry-II, Chemistry of d- and f- block elements

1. **Coordination Chemistry-II: (30L)**

VB description and its limitations. Elementary Crystal Field Theory: splitting of d^n configurations in octahedral, square planar and tetrahedral fields, crystal field stabilization energy (CFSE) in weak and strong fields; pairing energy. Spectrochemical series. Jahn- Teller distortion. Octahedral site stabilization energy (OSSE). Metal-ligand bonding (MO concept, elementary idea), sigma- and pi-bonding in octahedral complexes (qualitative pictorial approach) and their effects on the oxidation states of transitional metals (examples).

Magnetism and Colour: Orbital and spin magnetic moments, spin only moments of d^n ions and their correlation with effective magnetic moments, including orbital contribution; quenching of magnetic moment: super exchange and antiferromagnetic interactions (elementary idea with examples only); d-d transitions; L-S coupling; qualitative Orgel diagrams for $3d^1$ to $3d^9$ ions. Racah parameter. Selection rules for electronic spectral transitions; spectrochemical series of ligands; charge transfer spectra (elementary idea).

2 Chemistry of d- and f- block elements: (30L)

General comparison of 3d, 4d and 5d elements in term of electronic configuration, oxidation states, redox properties, coordination chemistry.

Lanthanoids and Actinoids:

General Comparison on Electronic configuration, oxidation states, colour, spectral and magnetic properties; lanthanide contraction, separation of lanthanides (ion-exchange method only).

Reference Books

1.Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006. 2. Greenwood,N.N. & Earnshaw A. Chemistry of the Elements, Butterworth-Heinemann. 1997.3. Cotton, F.A., Wilkinson, G., Murrillo, C. A., Bochmann, M., Advanced Inorganic Chemistry 6th Ed. 1999., Wiley. 4. Atkin, P. Shriver & Atkins' Inorganic Chemistry 5th Ed. Oxford University Press (2010). 5. Purecell, K.F. and Kotz, J.C., An Introduction to Inorganic Chemistry, Saunders: Philadelphia, 1980. 6. Sinha, S. P., Ed., Lanthanide and Actinide Research (Journal, Vol. 1, 1986). 7. Wulfsberg, G., Principles of Descriptive Inorganic Chemistry, Brooks/Cole: Monterey, CA, 1987.

CEMHP-11	**Inorganic Chemistry – IV**	**2 Credit**

Practical

Chromatography of metal ions

Principles involved in chromatographic separations. Paper chromatographic separation of following metal ions:

a) Ni (II) and Co (II)
b) Fe (III) and Al (III)

Gravimetry

a) Estimation of nickel (II) using Dimethylglyoxime (DMG).
b) Estimation of copper as CuSCN
c) Estimation of Al (III) by precipitating with oxine and weighing as Al(oxine)3 (aluminium oxinate)
d) Estimation of chloride

Spectrophotometry

a) Measurement of 10Dq by spectrophotometric method.
b) Determination of λ_{max} of $[Mn(acac)_3]$ and $[Fe(acac)_3]$ complexes

Reference Books

Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009.

CEMHT-12	**Organic Chemistry – V**	**4 Credit**

Theory: Carbocycles and Heterocycles, Cyclic Stereochemistry, Pericyclic reactions, Carbohydrates, Carbohydrates, Biomolecules

1. **Carbocycles and Heterocycles: (16L)**
   Polynuclear hydrocarbons and their derivatives: synthetic methods include Haworth, Bardhan-Sengupta, Bogert-Cook and other useful syntheses (with mechanistic details); fixation of double bonds and Fries rule; reactions (with mechanism) of naphthalene, anthracene, phenanthrene and their derivatives.

   Heterocyclic compounds: 5- and 6-membered rings with one heteroatom; reactivity, orientation and important reactions (with mechanism) of furan, pyrrole, thiophene and pyridine; synthesis (including retrosynthetic approach and mechanistic details): pyrrole: Knorr synthesis, Paal-Knorr synthesis, Hantzsch; furan: Paal-Knorr synthesis, Feist-Benary synthesis and its variation; thiophenes: Paal-Knorr synthesis, Hinsberg synthesis; pyridine: Hantzsch synthesis; benzo-fused 5- and 6-membered rings with one heteroatom: reactivity, orientation and important reactions (with mechanistic details) of indole, quinoline and isoquinoline; synthesis (including retrosynthetic approach and mechanistic details): indole: Fischer, Madelung and Reissert; quinoline: Skraup, Doebner- Miller, Friedlander; isoquinoline: Bischler-Napieralski synthesis.

2. **Cyclic Stereochemistry: (9L)**
   Alicyclic compounds: concept of I-strain; conformational analysis: cyclohexane, mono and disubstituted cyclohexane; symmetry properties and optical activity; topomerisation; ring-size and ease of cyclisation; conformation & reactivity in cyclohexane system: consideration of steric and stereoelectronic requirements; elimination (E2, E1), nucleophilic substitution (S_N1, S_N2, S_Ni, NGP), merged substitution-elimination; rearrangements; oxidation of cyclohexanol, esterification, saponification, lactonisation, epoxidation, pyrolytic syn elimination and fragmentation reactions.
3. **Pericyclic reactions: (9L)**
   Mechanism, stereochemistry, regioselectivity in case of Electrocyclic reactions: FMO approach involving 4π- and 6π-electrons (thermal and photochemical) and corresponding cycloreversion reactions.
   Cycloaddition reactions: FMO approach, Diels-Alder reaction, photochemical [2+2] cycloadditions.
   Sigmatropic reactions: FMO approach, sigmatropic shifts and their order; [1,3]- and [1,5]-H shifts and [3,3]-shifts with reference to Claisen and Cope rearrangements.

4. **Carbohydrates: (13L)**

Monosaccharides: Aldoses up to 6 carbons; structure of D-glucose & D- fructose (configuration & conformation); ring structure of monosaccharides (furanose and pyranose forms): Haworth representations and non-planar conformations; anomeric effect (including stereoelectronic explanation); mutarotation; epimerization; reactions (mechanisms in relevant cases): Fischer glycosidation, osazone formation, bromine-water oxidation, HNO_3 oxidation, selective oxidation of terminal $-CH_2OH$ of aldoses, reduction to alditols, Lobry de Bruyn-van Ekenstein rearrangement; stepping–up (Kiliani-Fischer method) and stepping–down (Ruff's & Wohl's methods) of aldoses; end-group-interchange of aldoses; acetonide (isopropylidene) and benzylidene protections; ring-size determination; Fischer's proof of configuration of (+)-glucose.

Disaccharides: Glycosidic linkages, concept of glycosidic bond formation by glycosyl donor-acceptor; structure of sucrose, inversion of cane sugar. **Polysaccharides:** starch (structure and its use as an indicator in titrimetric analysis).

5. **Biomolecules: (13L)**

Amino acids: synthesis with mechanistic details: Strecker, Gabriel, acetamido malonic ester, azlactone, Bücherer hydantoin synthesis, synthesis involving diketopiperazine; isoelectric point, zwitterions; electrophoresis, reaction (with mechanism): ninhydrin reaction, Dakin-West reaction; resolution of racemic amino acids.

Peptides: peptide linkage and its geometry; syntheses (with mechanistic details) of peptides using N-protection & C-protection, solid-phase (Merrifield) synthesis; peptide sequence: C-terminal and N-terminal unit determination (Edman, Sanger & 'dansyl' methods); partial hydrolysis; specific cleavage (enzymatic) of peptides: use of CNBr.

Nucleic acids: pyrimidine and purine bases (only structure & nomenclature); nucleosides and nucleotides corresponding to DNA and RNA; mechanism for acid catalysed hydrolysis of nucleosides (both pyrimidine and purine types); comparison of alkaline hydrolysis of DNA and RNA; elementary idea of double helical structure of DNA (Watson-Crick model); complimentary base–pairing in DNA.

Reference Books

1. Clayden, J., Greeves, N., Warren, S. Organic Chemistry, Second edition, Oxford University Press 2012. 2. Eliel, E. L. & Wilen, S. H. Stereochemistry of Organic Compounds, Wiley: London. 3. Nasipuri, D. Stereochemistry of Organic Compounds, Wiley Eastern Limited. 4. Sengupta, Subrata. Basic Stereochemistry of Organic molecules. 5. Kalsi, P. S. Stereochemistry

Conformation and Mechanism, Eighth edition, New Age International, 2014. 6. Fleming, I. Molecular Orbitals and Organic Chemical reactions, Reference/Student Edition, Wiley, 2009. 7. Fleming, I. Pericyclic Reactions, Oxford Chemistry Primer, Oxford University Press. 8. Gilchrist, T. L. & Storr, R. C. Organic Reactions and Orbital symmetry, Cambridge University Press. 9. Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 10. Finar, I. L. Organic Chemistry (Volume 2: Stereochemistry and the Chemistry of Natural Products), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 11. Morrison, R. T. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education). 12. Loudon, G. M. Organic Chemistry, Fourth edition, Oxford University Press. 13. Eames, J., Peach, J. M. Stereochemistry at a Glance, Blackwell Publishing, 2003. 14. Robinson, M. J. T., Stereochemistry, Oxford Chemistry Primer, Oxford University Press, 2005. 15. Davis, B. G., Fairbanks, A. J., Carbohydrate Chemistry, Oxford Chemistry Primer, Oxford University Press. 16. Joule, J. A. Mills, K. Heterocyclic Chemistry, Blackwell Science. 17. Acheson, R.M. Introduction to the Chemistry of Heterocyclic compounds, John Wiely & Sons (1976).18. Gilchrist, T. L. Heterocyclic Chemistry, 3rd edition, Pearson. 19. Bansal, R. K. Heterocyclic Chemistry, New Age International Publishers. 20. Davies, D. T., Heterocyclic Chemistry, Oxford Chemistry Primer, Oxford University Press.

CEMHP-12	**Organic Chemistry – V**	**2 Credit**

Practical

1. Column chromatographic separation of mixture of dyes
2. Paper chromatographic separation of a mixture containing 2/3 amino acids
3. Paper chromatographic separation of a mixture containing 2/3 sugars
4. TLC separation of a mixture containing 2/3 amino acids
5. TLC separation of a mixture of dyes (fluorescein and methylene blue)
6. Column chromatographic separation of leaf pigments from spinach leaves

Spectroscopic Analysis of Organic Compounds:

1. **Assignment** of labelled peaks in the 1H NMR spectra of the known organic compounds explaining the relative δ-values and splitting pattern.
2. **Assignment** of labelled peaks in the IR spectrum of the same compound explaining the relative frequencies of the absorptions (C-H, O-H, N-H, C-O, C-N, C-X, C=C, C=O, N=O, C≡C, C≡N stretching frequencies; characteristic bending vibrations are included).
3. The students must record full spectral analysis of at least 15 (fifteen) compounds from the following list:

1) 4-Bromoacetanilide
2) Vanillin
3) 4-Aminobenzoic acid

4) Salicylamide
5) 1,3-Dinitrobenzene
6) Benzylacetate
7) 4-Nitrobenzaldehyde
8) Mesityl oxide
9) 2-Hydroxybenzaldehyde
10) 4-Nitroaniline
11) 3-Nitrobenzaldehyde
12) 2-Methoxybenzaldehyde
13) Methyl 4-hydroxybenzoate
14) 3-Aminobenzoic acid
15) 3-Methylacetanilide
16) 5-Methyl-2-nitroanisole
17) 2-Bromo-4'-methylacetophenone
18) 2-Methoxyacetophenone
19) 2-Hydroxyacetophenone
20) trans-4-NitrocinnamaldehydE
21) Diethyl fumarate
22) 4-Methylacetanilide
23) 2-Hydroxy-3-nitrobenzaldehyde
24) 2,3-Dimethylbenzonitrile
25) Pent-1-yn-3-ol
26) 3-Ethoxy-4-hydroxybenzaldehyde
27) Methyl 3-hydroxybenzoate
28) Ethyl 3-aminobenzoate
29) Ethyl 4-aminobenzoate
30) 3-nitroanisole

Reference Books

1.University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N. University of Calcutta, 2003. 2. Practical Workbook Chemistry (Honours), UGBS, Chemistry, University of Calcutta, 2015. 3. Furniss, B.S.; Hannaford, A.J.; Smith, P.W.G.; Tatchell, A.R. Practical Organic Chemistry, 5th Ed., Pearson (2012). 4. Mann, F.G. & Saunders, B.C. Practical Organic Chemistry, Pearson Education.

CEMHTDSE-1A	**Advanced Physical Chemistry**	**4 Credit**

Theory: Crystal Structure, Statistical Thermodynamics, Special selected topics,

1. **Crystal Structure** **(20 L)**

Bravais Lattice and Laws of Crystallography: Types of solid, Bragg's law of diffraction; Laws of crystallography (Haüy's law and Steno's law); Permissible symmetry axes in crystals; Lattice, space lattice, unit cell, crystal planes, Bravais lattice. Packing of uniform hard

sphere, close packed arrangements (fcc and hcp).
Tetrahedral and octahedral voids Void space in P-type,F-type, I-type cubic systems. Crystal planes: Distance between consecutive planes [cubic, tetragonal and orthorhombic lattices]; Indexing of planes, Miller indices; calculation of dhkl; Relation between molar mass and unit cell dimension for cubic system; Bragg's law (derivation).
Determination of crystal structure: Powder method; Structure of NaCl and KCl crystals.

2. **Statistical Thermodynamics (18L)**

Configuration: Macrostates, microstates and configuration; calculation with harmonic oscillator, variation of W with E; equilibrium configuration.
Boltzmann distribution: Thermodynamic probability, entropy and probability, Boltzmann distribution formula (with derivation); Applications to barometric distribution; Partition function, concept of ensemble -canonical ensemble and grand canonical ensembles.
Partition function: molecular partition function and thermodynamic properties, Maxwell's speed distribution, Gibbs' paradox.

3. **Special selected topics (22 L)**

Specific heat of solid: Coefficient of thermal expansion, thermal compressibility of solids; Dulong –Petit's law; Perfect Crystal model, Einstein's theory – derivation from partition function, limitations. Debye's T^3 law - analysis at two extremes.
3rd law: Absolute entropy, Plank's law, Calculation of entropy, Nernst heat theorem. Adiabatic demagnetization approach to zero Kelvin, adiabatic cooling, demagnetization, adiabatic demagnetization, involved curves.
Polymers: Classification of polymers, nomenclature, Molecular forces and chemical bonding in polymers, Texture of Polymers; Criteria for synthetic polymer formation; Relationships between functionality, extent of reaction and degree of polymerization. Mechanism and kinetics of step growth & copolymerization, conducting polymers.

Reference Books

1.Castellan, G. W. Physical Chemistry, Narosa. 2. Levine, I. N. Physical Chemistry, Tata McGraw-Hill. 3. Moore, W. J. Physical Chemistry, Orient Longman. 4. Atkins, P. W. & Paula, J. de Atkins', Physical Chemistry, Oxford University Press. 5. McQuarrie, D. A. & Simons, J. D. Physical Chemistry: A Molecular Approach, Viva Press. 6. Engel, T. & Reid, P. Physical Chemistry, Pearson. 7. Nash, L. K. Elements of Statistical Thermodynamics, Dover. 8. Rastogi, R. P. & Misra, R.R. An Introduction to Chemical Thermodynamics, Vikas. 9. Zemansky, M. W. & Dittman, R.H. Heat

CEMHPDSE-1A	**Advanced Physical Chemistry**	**2 Credit**

Practical

Computer Programming based on numerical methods for:

1) Roots of equations: (e.g. volume of van der Waals gas and comparison with ideal gas, pH of a weak acid)
2) Numerical differentiation (e.g., change in pressure for small change in volume of a van der Waals gas, potentiometric titrations)
3) Numerical integration (e.g. entropy/ enthalpy change from heat capacity data), probability distributions (gas kinetic theory) and mean values
4) Matrix operations (Application of Gauss-Siedel method in colourimetry)
5) Simple exercises using molecular visualization software

Reference Books

1. McQuarrie, D. A. Mathematics for Physical Chemistry University Science Books (2008). 2. Mortimer, R. Mathematics for Physical Chemistry. 3rd Ed. Elsevier (2005). 3. Yates, P. Chemical Calculations. 2nd Ed. CRC Press (2007). 4. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman (2007) Chapters 3-5. 5. Noggle, J. H. Physical Chemistry on a Microcomputer. Little Brown & Co. (1985).

CEMHTDSE-2A	**Analytical Methods in Chemistry**	**4 Credit**

Theory: Qualitative and quantitative, Optical methods of analysis. Thermal and Electroanalytical methods of analysis. Separation techniques

1. Qualitative and quantitative aspects of analysis (10L)

Sampling, evaluation of analytical data, errors, accuracy and precision, methods of their expression, normal law of distribution of indeterminate errors, statistical test of data; F, Q and t test, rejection of data, and confidence intervals

2. Optical methods of analysis (20L)

i) Origin of spectra, interaction of radiation with matter, fundamental laws of spectroscopy and selection rules, validity of Beer-Lambert's law.

ii) **UV-Visible Spectrometry:** Basic principles of instrumentation (choice of source, monochromator and detector) for single and double beam instrument;

iii) **Basic principles of quantitative analysis:** estimation of metal ions from aqueous solution, geometrical isomers, keto-enol tautomers. Determination of composition of metal complexes using Job's method of continuous variation and mole ratio method.

iv) **Infrared Spectrometry:** Basic principles of instrumentation (choice of source, monochromator & detector) for single and double beam instrument; sampling techniques. Structural illustration through

interpretation of data, Effect and importance of isotope substitution.

v) **Flame Atomic Absorption and Emission Spectrometry:** Basic principles of instrumentation (choice of source, monochromator, and detector, choice of flame and Burner designs. Techniques of atomization and sample introduction; Method of background correction, sources of chemical interferences and their method of removal. Techniques for the quantitative estimation of trace level of metal ions from water samples.

vi) **Thermal methods of analysis:** Theory of thermogravimetry (TG), instrumentation. Composition determination of Ca and Mg from their mixture.

4. Electroanalytical methods (10L)

Classification of electroanalytical methods, basic principle of pH metric, potentiometric and conductometric titrations. Techniques used for the determination of equivalence points. Techniques used for the determination of pKa values.

5. Separation techniques (12L)

i. **Solvent extraction:** Classification, principle and efficiency of the technique. Mechanism of extraction: extraction by solvation and chelation.

ii. **Technique of extraction:** batch, continuous and counter current extractions.

iii. **Qualitative and quantitative aspects of solvent extraction**: extraction of metal ions from aqueous solution, extraction of organic species from the aqueous and nonaqueous media.

iv. **Chromatography:** Classification, principle and efficiency of the technique. Mechanism of separation: adsorption, partition & ion exchange.

v. **Development of chromatograms:** frontal, elution and displacement methods

vi. **Qualitative and quantitative aspects of chromatographic methods of analysis:** IC, GLC, GPC, TLC and HPLC.

vii. Separation and analysis using GC and HPLC (dye and pesticide analysis)

viii. Role of computers in instrumental methods of analysis

Reference Books

1. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009.2. Willard, H.H. Et al.: Instrumental Methods of Analysis, 7th Ed. Wardsworth Publishing Company, Belmont, California, USA, 1988. 3. Christian, G.D. Analytical Chemistry, 6th Ed. John Wiley & Sons, New York, 2004. 4. Harris, D.C.: Exploring Chemical Analysis, 9th Ed. New York, W.H. Freeman, 2016. 5. Khopkar, S.M. Basic Concepts of Analytical Chemistry. New Age International Publisher, 2009. 6. Skoog, D.A. Holler F.J. & Nieman, T.A. Principles of Instrumental Analysis, Cengage Learning India Ed. 7. Mikes, O. Laboratory Hand Book of Chromatographic & Allied Methods, Elles Harwood Series on Analytical Chemistry, John Wiley & Sons, 1979. 8. Ditts, R.V. Analytical Chemistry; Methods of separation, van Nostrand, 1974.

CEMHTDSE- 2B	**Polymer Chemistry**	**4Credit**

Theory: Introduction, Functionality and its importance, Kinetics of Polymerization, Crystallization and crystallinity, Nature and structure of polymers, molecular weight of polymers, glass transition temperature (Tg) , polymer solution, Properties of polymer

1. **Introduction and history of polymeric materials (4L)**
   Different schemes of classification of polymers, Polymer nomenclature, Molecular forces and chemical bonding in polymers, Texture of Polymers.
2. **Functionality and its importance (6L)**
   Criteria for synthetic polymer formation, classification of polymerization processes, elationships between functionality, extent of reaction and degree of polymerization. Bi-functional systems, Poly- functional systems.
3. **Kinetics of Polymerization (8L)**
   Mechanism and kinetics of step growth, radical chain growth, ionic chain (both cationic and anionic) and coordination polymerizations.
4. **Crystallization and crystallinity (4L)**
   Determination of crystalline melting point and degree of crystallinity, Morphology of crystalline polymers, Factors affecting crystalline melting point.
5. **Nature and structure of polymers (4L)**
   Structure Property relationships.
6. **Determination of molecular weight of polymers (6L)**
   (M_n, M_w, etc) by end group analysis, viscometry, light scattering and osmotic pressure methods.Molecular weight distribution and its significance. Polydispersity index.
7. **Glass transition temperature (Tg) and determination of Tg (4L)**
8. Free volume theory, WLF equation, Factors affecting glass transition temperature (Tg).
9. **Polymer Solution (10L)**
   Criteria for polymer solubility, Solubility parameter, Thermodynamics of polymer solutions, entropy, enthalpy, and free energy change of mixing of polymers solutions, Flory-Huggins theory , Lower and Upper critical solution temperatures.

9. **Properties of Polymer (14L)**
   (Physical, thermal, Flow & Mechanical Properties)
   Brief introduction to preparation, structure, properties and application of the following polymers:
   Polyolefins, polystyrene and styrene copolymers, poly(vinyl chloride) and related polymers, poly(vinyl acetate) and related polymers, acrylic polymers, fluoro polymers, Polyamides and related polymers. Phenol formaldehyde resins (Bakelite, Novalac), polyurethanes, silicone polymers, polydienes, Polycarbonates, Conducting Polymers,

[polyacetylene, polyaniline, poly(p-phenylene sulphide polypyrrole, polythiophene)].

Reference Book

1.R.B. Seymour& C.E. Carraher: Polymer Chemistry: An Introduction, Marcel Dekker, Inc. New York, 1981. 2. G. Odian: Principles of Polymerization, 4th Ed. Wiley, 2004. F.W. Billmeyer: Textbook of Polymer Science, 2nd Ed. Wiley Interscience, 1971. P. Ghosh: Polymer Science & Technology, Tata McGraw-Hill Education, 1991. 5. R.W. Lenz: Organic Chemistry of Synthetic High Polymers. Interscience Publishers, New York, 1967.

CEMHPDSE-2A	**Analytical Methods in Chemistry**	**2 Credit**

Practical

Separation Techniques – Chromatography

1) **Separation of mixtures:** Separation and identification of the monosaccharides present in the given mixture (glucose & fructose) by paper chromatography. Reporting the RF values.

2) Separate a mixture of Sudan yellow and Sudan Red by TLC technique and identify them on the basis of their RF values.
3) Chromatographic separation of the active ingredients of plants, flowers and juices by TLC

Solvent Extractions

1) To separate a mixture of Ni^{2+} & Fe^{2+} by complexation with DMG and extracting the Ni^{2+}- DMG complex in chloroform, and determine its concentration by spectrophotometry.

2) Analysis of soil:
   a) Determination of pH of soil.
   b) Total soluble salt
   c) Estimation of calcium, magnesium, phosphate, nitrate

3) Ion exchange:
   a) Determination of exchange capacity of cation exchange resins and anion exchange resins.

Spectrophotometry

1) Determination of pK_a values of indicator using spectrophotometry
2) Determination of chemical oxygen demand (COD)
3) Determination of Biological oxygen demand (BOD)

Reference Books

1. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009. 2. Willard,H.H. Et al.: Instrumental Methods of Analysis, 7th Ed. Wardsworth Publishing Company, Belmont, California, USA, 1988. 3. Christian, G.D. Analytical Chemistry, 6th Ed. John Wiley & Sons, New York, 2004. 4. Harris, D.C.: Exploring Chemical Analysis, 9th Ed. New York,W.H. Freeman, 2016. 5. Khopkar, S.M. Basic Concepts of Analytical Chemistry. New Age International Publisher, 2009. 6. Skoog, D.A. Holler F.J. & Nieman, T.A. Principles of Instrumental Analysis, Cengage Learning India **Ed. 7. Mikes,** O.

Laboratory Hand Book of Chromatographic & Allied Methods, Elles Harwood Series on Analytical Chemistry, John Wiley & Sons, 1979. 8. Ditts, R.V. Analytical Chemistry; Methods of separation, van Nostrand, 1974.

CEMHPDSE-2B	**Polymer Chemistry**	**2 Credit**

Practical

1. **Polymer Synthesis**
   a. Free radical solution polymerization of styrene(St)/Methyl Methacrylate (MMA) / Methyl Acrylate (MA) / Acrylic acid (AA).
   b. Polymerization using benzoyl peroxide (BPO) / 2,2'-azo-bis-isobutylonitrile (AIBN)
   c. Preparation of nylon 66/6.
2. **Polymer characterization**
   a. Determination of molecular weight by viscometry:
   i. Polyacrylamide-aq.$NaNO_2$ solution
   ii. Poly vinyl proplylidine (PVP) in water
   b. Determination of the viscosity-average molecular weight of poly(vinyl alcohol) (PVOH) and the fraction of "head-to-head" monomer linkages in the polymer.
   c. Determination of molecular weight by end group analysis: Polyethylene glycol (PEG) (OH group).
   d. Testing of mechanical properties of polymers
   e Determination of hydroxyl number of a polymer using colorimetric method.

3. **Polymer analysis**

i)Estimation of the amount of HCHO in the given solution by sodium sulphite method.

ii) IR studies of polymers

Reference Books

1.M.P. Stevens, Polymer Chemistry: An Introduction, 3rd Ed., Oxford University Press, 1999. 2. H.R. Allcock, F.W. Lampe & J.E. Mark, Contemporary Polymer Chemistry, 3rd ed. Prentice-Hall (2003). 3. F.W. Billmeyer, Textbook of Polymer Science, 3rd ed. Wiley-Interscience (1984) . 4. J.R. Fried, Polymer Science and Technology, 2nd ed. Prentice-Hall (2003). 5. P. Munk & T.M. Aminabhavi, Introduction to Macromolecular Science, 2nd ed. John Wiley & Sons (2002). 6. L. H. Sperling, Introduction to Physical Polymer Science, 4th ed. John Wiley & Sons (2005). 7. M.P. Stevens, Polymer Chemistry: An Introduction 3rd ed. Oxford University Press (2005). 8. Seymour/ Carraher's Polymer Chemistry, 9th ed. by Charles E. Carraher, Jr. (2013).

CEMHSE-1A	IT skills for Chemist	02 Credit

1. Mathematics (10L)

i. Fundamentals, mathematical functions, polynomial expressions, logarithms, the exponential function, units of a measurement, interconversion of units, constants and variables, equation of a straight line, plotting graphs.

ii. Uncertainty in experimental techniques: Displaying uncertainties, measurements in chemistry, decimal places, significant figures, combining quantities.

iii. Uncertainty in measurement: types of uncertainties, combining uncertainties. Statistical treatment. Mean, standard deviation, relative error. Data reduction and the propagation of errors. Graphical and numerical data reduction. Numerical curve fitting: the method of least squares (regression).

iv. Algebraic operations on real scalar variables (e.g. manipulation of van der Waals equation in different forms). Roots of quadratic equations analytically and iteratively (e.g. pH of a weak acid). Numerical methods of finding roots (Newton- Raphson, binary –bisection, e.g. pH of a weak acid not ignoring the ionization of water, volume of a van der Waals gas, equilibrium constant expressions).

v. Differential calculus: The tangent line and the derivative of a function, numerical differentiation (e.g., change in pressure for small change in volume of a van der Waals gas, potentiometric titrations).

vi. Numerical integration (Trapezoidal and Simpson's rule, e.g. entropy/enthalpy change from heat capacity data).

2. **Computer programming (10L)**

Constants, variables, bits, bytes, binary and ASCII formats, arithmetic expressions, hierarchy of operations, inbuilt functions. Elements of the BASIC language. BASIC keywords and commands. Logical and relative operators. Strings and graphics. Compiled versus interpreted languages. Debugging. Simple programs using these concepts. Matrix addition and multiplication. Statistical analysis. BASIC programs for curve fitting, numerical differentiation and integration (Trapezoidal rule, Simpson's rule), finding roots (quadratic formula, iterative, Newton-Raphson method).

3. **Hands On (10L)**

i. Introductory writing activities: Introduction to word processor and structure drawing (ChemSketch) software. Incorporating chemical structures, chemical equations, and expressions from chemistry (e.g. Maxwell-Boltzmann distribution law, Bragg's law, van der Waals equation, etc.) into word processing documents.

ii. **Handling numeric data:** Spreadsheet software (Excel), creating a spreadsheet, entering and formatting information, basic functions and formulae, creating charts, tables and graphs. Incorporating tables and

graphs into word processing documents. Simple calculations, plotting graphs using a spreadsheet (Planck's distribution law, radial distribution curves for hydrogenic orbitals, gas kinetic theory- Maxwell-Boltzmann distribution curves as function of temperature and molecular weight), spectral data, pressure-volume curves of van der Waals gas (van der Waals isotherms), data from phase equilibria studies. Graphical solution of equations.

iii. **Numeric modelling:** Simulation of pH metric titration curves. Excel functions LINEST and Least Squares. Numerical curve fitting, linear regression (rate constants from concentration- time data, molar extinction coefficients from absorbance data), numerical differentiation (e.g. handling data from potentiometric and pH metric titrations, pKa of weak acid), integration (e.g. entropy/enthalpy change from heat capacity data).

iv. **Statistical analysis:** Gaussian distribution and Errors in measurements and their effect on data sets. Descriptive statistics using Excel. Statistical significance testing: The t test. The F test.

v. **Presentation:** Presentation graphics

Reference Books

1. McQuarrie, D. A. Mathematics for Physical Chemistry University Science Books (2008). 2. Mortimer, R. Mathematics for Physical Chemistry. 3rd Ed. Elsevier (2005). 3. Steiner, E. The Chemical Maths Book Oxford University Press (1996). 4. Yates, P. Chemical calculations. 2nd Ed. CRC Press (2007). 5. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman (2007) Chapters 3-5. 6. Levie, R. de, How to use Excel in analytical chemistry and in general scientific data analysis, Cambridge Univ. Press (2001) 487 pages. 7. Noggle, J. H. Physical chemistry on a Microcomputer. Little Brown & Co. (1985). 8. Venit, S.M. Programming in BASIC: Problem solving with structure and style. Jaico Publishing House: Delhi (1996).

CEMHSE-1B	**Basic Analytical Chemistry**	**02 Credit**

1. **Introduction** **(2L)**

Strategies of Analytical Chemistry and its interdisciplinary applicability. Protocol of sampling. Variability and validity of analytical measurements. Presentation of experimental data and results, from the point of view of significant figures.

Complexometry

Complexometric titrations, Chelation, Chelating agents, use of indicators. Estimation of Calcium and Magnesium ions as Calcium carbonate by complexometric titration.

Soil Analysis

Composition, pH of soil samples, estimation of calcium and magnesium content.

2. **Analysis of water** **(4L)**

Definition of pure water, sources responsible for contaminating water, water sampling methods, water purification methods.

Determination of pH, acidity and alkalinity of a water sample. Determination of Biological Oxygen Demand (BOD).

3. **Analysis of food products (4L)**

Nutritional value of foods, idea about food processing and food preservations and adulteration.

Identification of adulterants in some common food items like coffee powder, asafoetida, chilli powder, turmeric powder, coriander powder and pulses, etc.

Analysis of preservatives and colouring matter.

4. **Chromatography (4L)**

Definition, general introduction on principles of chromatography,

paper chromatography, TLC etc. Paper chromatographic separation of mixture of metal ion (Fe3+ and Al3+). To compare paint samples

by TLC method.

5. **Ion-exchange (4L)**

Column, ion-exchange chromatography etc. Determination of ion exchange capacity of anion / cation exchange resin (using batch procedure if use of column is not feasible).

6. **Analysis of cosmetics (3L)**

Major and minor constituents and their function

Analysis of deodorants and antiperspirants, Al, Zn, boric acid, chloride, sulphate. Determination of constituents of talcum powder: Magnesium oxide, Calcium oxide, Zinc oxide and Calcium carbonate by complexometric titration

7. **Suggested Applications (Any one) (2L)**

To study the use of phenolphthalein in trap cases. To analyse arson accelerants.

To carry out analysis of gasoline.

8. **Suggested Instrumental demonstrations (3L)**

Estimation of macro nutrients: Potassium, Calcium, Magnesium in soil samples by flame photometry.

Spectrophotometric determination of Iron in Vitamin / Dietary Tablets. Spectrophotometric Identification and Determination of Caffeine and Benzoic Acid in Soft Drink

Reference Books

1. Willard, H.H., Merritt, L.L., Dean, J. & Settoe, F.A. Instrumental Methods of Analysis, 7th Ed. Wadsworth Publishing Company Ltd., Belmont, California, USA, 1988. 2. Skoog, D.A., Holler, F.J. & Crouch, S. Principles of Instrumental Analysis, Cengage Learning India Edition, 2007. 3. Skoog, D.A.; West, D.M. & Holler, F.J. Analytical Chemistry: An Introduction 6th Ed., Saunders College Publishing, Fort Worth, Philadelphia (1994). 4. Harris, D. C. Quantitative Chemical Analysis, 9th ed. Macmillan Education, 2016. 5. Dean, J. A. Analytical Chemistry

Handbook, McGraw Hill, 2004. 6. Day, R. A. & Underwood, A. L. Quantitative Analysis, Prentice Hall of India, 1992. 7. Freifelder, D.M. Physical Biochemistry 2nd Ed., W.H. Freeman & Co.,N.Y. USA (1982). 8. Cooper, T.G. The Tools of Biochemistry, John Wiley & Sons,N.Y. USA. 16 (1977). 9. Vogel, A. I. Vogel's Qualitative Inorganic Analysis 7th Ed., Prentice Hall, 1996. 10. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009. 11. Robinson, J.W. Undergraduate Instrumental Analysis 5th Ed., Marcel Dekker, Inc., New York (1995). 12. Christian, G.D. Analytical Chemistry, 6th Ed. John Wiley & Sons, New York, 2004

Semester - VI

CEMHT-13	**Inorganic Chemistry – V**	**4 Credit**

Theory: Bioinorganic Chemistry; Organometallic Chemistry and Catalysis: Reaction Kinetics and Mechanism

1. **Bioinorganic Chemistry: (25L)**

Elements of life: essential and beneficial elements, major, trace and ultratrace elements. Basic chemical reactions in the biological systems and the role of metal ions (specially Na^+, K^+, Mg^{2+},Ca^{2+}, $Fe^{3+/2+}$, $Cu^{2+/+}$, and Zn^{2+}). Metal ion transport across biological membrane Na^+/ K^+-ion pump.
Dioxygen molecule in life.
Dioxygen management proteins: Haemoglobin, Myoglobin, Hemocyanine and Hemerythrin.
Electron transfer proteins: Cytochromes and Ferredoxins.
Hydrlytic enzymes: carbonate bicarbonate buffering system and carbonic anhydrase and carboxyanhydrase A.
Biological nitrogen fixation, Photosynthesis: Photosystem-I and Photosystem-II. Toxic metal ions and their effects, chelation therapy (examples only), Pt and Au complexes as drugs (examples only), metal dependent diseases (examples only).

2. **Organometallic Chemistry and Catalysis: (25L)**

Definition and classification of organometallic compounds on the basis of bond type. Concept of hapticity of organic ligands. 18-electron and 16-electron rules (pictorial MO approach). Applications of 18-electron rule to metal carbonyls, nitrosyls, cyanides. General methods of preparation of mono and binuclear carbonyls of 3d- series. Structures of mononuclear and binuclear carbonyls. pi-acceptor behaviour of CO, synergic effect and use of IR data to explain extent of back bonding. Zeise's salt: Preparation, structure, evidences of synergic effect. Ferrocene: Preparation and reactions (acetylation, alkylation, metallation, Mannich Condensation).: substitution, oxidative addition, reductive elimination and insertion reactions.
Study of the following industrial processes
1. Alkene hydrogenation (Wilkinson's Catalyst)
2. Hydroformylation

3. Wacker Process
4. Synthetic gasoline (Fischer Tropsch reaction)
5. Ziegler-Natta catalysis for olefin polymerization.

3. Reaction Kinetics and Mechanism: (10L)

Introduction to inorganic reaction mechanisms. Substitution reactions in square planar complexes, Trans- effect and its application in complex synthesis, theories of trans effect, Mechanism of nucleophilic substitution in square planar complexes, Thermodynamic and Kinetic stability, Kinetics of octahedral substitution, Ligand field effects and reaction rates, Mechanism of substitution in octahedral complexes.

Reference Books

1.Lippard, S.J. & Berg, J.M. Principles of Bioinorganic Chemistry Panima Publishing Company 1994. 2. Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006. 3.Greenwood, N.N. & Earnshaw A. Chemistry of the Elements, Butterworth- Heinemann, 1997. 4.Cotton, F.A., Wilkinson, G., Murrillo, C. A., Bochmann, M., Advanced Inorganic Chemistry 6th Ed. 1999., Wiley. 5.Bertini, I., Gray, H. B., Lippard, S.J., Valentine, J. S., Viva, 2007. 6.Basolo, F, and Pearson, R.C. Mechanisms of Inorganic Chemistry, John Wiley & Sons, NY, 1967. 7. Purecell, K.F. and Kotz, J.C., An Introduction to Inorganic Chemistry, Saunders: Philadelphia, 1980. 8. Powell, P. Principles of Organometallic Chemistry, Chapman and Hall, 1988. 9.Collman, J. P. et al. Principles and Applications of Organotransition Metal Chemistry. Mill Valley, CA: University Science Books, 1987. 10.Crabtree, R. H. The Organometallic Chemistry of the Transition Metals. New York, NY: John Wiley, 2000.

CEMHP-13	**Inorganic Chemistry – V**	**2 Credit**

Practical

Qualitative semimicro analysis

Qualitative semimicro analysis of mixtures containing four radicals. Emphasis should be given to the understanding of the chemistry of different reactions and to assign the most probable composition.

Cation Radicals: Na^{+}, K^{+}, Ca^{2+}, Sr^{2+}, Ba^{2+}, Al^{3+}, Cr^{3+}, Mn^{2+}/Mn^{4+}, Fe^{3+}, Co^{2+}/Co^{3+}, Ni^{2+}, Cu^{2+}, Zn^{2+}, Pb^{2+}, Cd^{2+}, Bi^{3+}, Sn^{2+}/Sn^{4+}, As^{3+}/As^{5+}, $Sb^{3+/5+}$, NH^{+}, Mg^{2+}.

Anion Radicals: F^{-}, Cl^{-}, Br^{-}, BrO_3^{-}, I^{-}, IO_3^{-}, SCN^{-}, S^{2-}, SO_4^{2-}, NO_3^{-}, NO_2^{-}, PO_4^{3-}, AsO_4^{3-}, BO_3^{3-}, CrO_4^{2-} / $Cr_2O_7^{2-}$, $Fe(CN)_6^{4-}$, $Fe(CN)_6^{3-}$.

Insoluble Materials: Al_2O_3(ig), Fe_2O_3(ig), Cr_2O_3(ig), SnO_2, $SrSO_4$, $BaSO_4$, CaF_2, $PbSO_4$.

Reference Books

Svehla, G., Vogel's Qualitative Inorganic Analysis, Pearson Education, 2012.

CEMHT-14	**Physical Chemistry – IV**	**4 Credit**

Molecular Spectroscopy, Photochemistry, Surface Phenomenon

1. Molecular Spectroscopy (24 L)

Interaction of electromagnetic radiation with molecules and various types of spectra; Born-Oppenheimer approximation

Rotation spectroscopy: Selection rules, intensities of spectral lines, determination of bond lengths of diatomic and linear triatomic molecules, isotopic substitution.

Vibrational spectroscopy: Classical equation of vibration, computation of force constant, amplitude of diatomic molecular vibrations, anharmonicity, Morse potential, dissociation energies, fundamental frequencies, overtones, hot bands, degrees of freedom for polyatomic molecules, modes of vibration, concept of group frequencies; Diatomic vibrating rotator, P, Q, R branches.

Raman spectroscopy: Qualitative treatment of Rotational Raman effect; Vibrational Raman spectra, Stokes and anti-Stokes lines. Effect of nuclear spin, their intensity difference, rule of mutual exclusion.

Nuclear Magnetic Resonance (NMR) spectroscopy: Principles of NMR spectroscopy, Larmor precession, chemical shift and low resolution spectra, different scales, spin-spin coupling and high resolution spectra, interpretation of PMR spectra of organic molecules.

Electron Spin Resonance (ESR) spectroscopy: Its principle, Hyperfine structure, ESR of simple radicals.

2. Photochemistry (18 L)

Lambert-Beer's law: Characteristics of electromagnetic radiation, Lambert- Beer's law and its limitations, physical significance of absorption coefficients; Laws of photochemistry, Stark-Einstein law of photochemical equivalence quantum yield, actinometry, examples of low and high quantum yields.

Photochemical Processes: Potential energy curves (diatomic molecules), Frank- Condon principle and vibrational structure of electronic spectra; Bond dissociation and principle of determination of dissociation energy (ground state); Decay of excited states by radiative and non-radiative paths; Pre-dissociation; Fluorescence and phosphorescence, Jablonskii diagram.

Rate of Photochemical processes: Photochemical equilibrium and the differential rate of photochemical reactions, Photostationary state; HI decomposition, H_2-Br_2 reaction, dimerisation of anthracene; photosensitised reactions, quenching; Role of photochemical reactions in

biochemical processes, photostationary states, chemiluminescence.

3. Surface phenomenon (18 L)

Surface tension and energy: Surface tension, surface energy, excess pressure, capillary rise and surface tension; Work of cohesion and adhesion, spreading of liquid over other surface; Vapour pressure over curved surface; Temperature dependence of surface tension.

Adsorption: Physical and chemical adsorption; Freundlich and Langmuir adsorption isotherms; multilayer adsorption and BET isotherm (no derivation required); Gibbs adsorption isotherm and surface excess; Heterogenous catalysis (single reactant); Zero order and fractional order reactions.

Colloids: Lyophobic and lyophilic sols, Origin of charge and stability of lyophobic colloids, coagulation and Schultz-Hardy rule, Zeta potential and Stern double layer (qualitative idea), Tyndall effect; Electrokinetic phenomena (qualitative idea only); Determination of Avogadro number by Perrin's method; Stability of colloids and zeta potential; Micelle formation.

Reference Books

1. Castellan, G. W. Physical Chemistry, Narosa. 2. Levine, I. N. Physical Chemistry, Tata McGraw-Hill . 3. Atkins, P. W. & Paula, J. de Atkin's, Physical Chemistry, Oxford University Press. 4. McQuarrie, D. A. & Simons, J. D. Physical Chemistry: A Molecular Approach, Viva Press. 5. Mortimer, R. G. Physical Chemistry, Elsevier. 6. Laidler, K. J.Chemical Kinetics, Pearson. 7. Banwell, C. N. Fundamentals of Molecular Spectroscopy, Tata-McGraw-Hill. 8. Barrow, G. M. Molecular Spectroscopy, McGraw- Hill. 9. Hollas, J.M. Modern Spectroscopy, Wiley India. 10. McHale, J. L. Molecular Spectroscopy, Pearson Education. 11. Wayne, C. E. & Wayne, R. P. Photochemistry, OUP. 12. Brown, J. M. Molecular Spectroscopy, OUP. 13. Levine, I. N. Quantum Chemistry, PHI. 14. Atkins, P. W. Molecular Quantum Mechanics, Oxford.

CEMHP-14	**Physical Chemistry – IV**	**2 Credit**

Practical

1. Determination of surface tension of a liquid using Stalagmometer.
2. Determination of CMC from surface tension measurements.
3. Verification of Beer and Lambert's Law for $KMnO_4$ and $K_2Cr_2O_7$ solution.
4. Study of kinetics of $K_2S_2O_8$+ KI reaction, spectrophotometrically.
5. Determination of pH of unknown buffer, spectrophotometrically.
6. Spectrophotometric determination of CMC.

Reference Books

Viswanathan, B., Raghavan, P.S. Practical Physical Chemistry Viva Books (2009). Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson. 3. Harris, D. C. Quantitative Chemical Analysis. 6th Ed., Freeman

(2007.). 4. Palit, S.R., De, S.K. Practical Physical Chemistry Science Book Agency. 5. University Hand Book of Undergraduate Chemistry Experiments, edited by Mukherjee, G. N., University of Calcutta. 6. Levitt, B. P. edited Findlay's Practical Physical Chemistry Longman Group Ltd. 7. Gurtu, J. N., Kapoor, R., Advanced Experimental Chemistry S. Chand & Co. Ltd

CEMHTDSE-3A	**Instrumental Methods of Chemical Analysis**	**4 Credit**

Theory: Instrumental Methods of Chemical Analysis

1. **Introduction to spectroscopic methods of analysis (8L)**

Recap of the spectroscopic methods covered in detail in the core chemistry syllabus: Classification of analytical methods and the types of instrumental methods. Consideration of electromagnetic radiation

2. **Molecular spectroscopy (14L)**

Infrared spectroscopy: Interactions with molecules: absorption and scattering. Means of excitation (light sources), separation of spectrum (wavelength dispersion, time resolution), detection of the signal (heat, differential detection), interpretation of spectrum (qualitative, mixtures, resolution), advantages of Fourier Transform (FTIR). Samples and results expected. Applications: Issues of quality assurance and quality control, Special problems for portable instrumentation and rapid detection.

UV-Visible/ Near IR – emission, absorption and fluorescence: Excitation sources (lasers, time resolution), wavelength dispersion (grating, prism, filter). Resolution, detection of signal (photocells, photomultipliers, diode arrays), sensitivity and S/N, Single and Double Beam instruments, Interpretation (quantification, mixtures).

3. **Chromatography (12L)**

Principles of Gas chromatography, liquid chromatography, supercritical fluid chromatography, Importance of column chromatographic technology (packing, capillaries), Separation based on increasing number of factors (volatility, solubility, interactions with stationary phase, size, electrical field)

Detection of different samples, single and coupled / hyphenated detector

4. **Elemental analysis (12L)**

Mass spectrometry (electrical discharges).

Atomic spectroscopy: Atomic absorption, Atomic emission, and Atomic fluorescence.

Excitation and atomisation (flames, electrical discharges, plasmas), Wavelength separation and resolution (dependence on technique), Detection of radiation (simultaneous/scanning, signal noise), Interpretation (errors due to molecular and ionic species, matrix effects, other interferences).

5. **NMR spectroscopy (6L)**

Principle, Instrumentation, Factors affecting chemical shift, Spin- coupling,

Applications.

6. **Electroanalytical techniques (4L)**
   Potentiometry & Voltammetry

7. **Radiochemical Methods: Elementary idea (2L)**
8. **Basic idea of X-ray analysis and electron spectroscopy (surface analysis) (2L)**

Reference Books

1.D.A. Skoog, F.J. Holler & S. Crouch (ISBN 0-495-01201-7) Principles of Instrumental Analysis, Cengage Learning India Edition, 2007. 2. Willard, Merritt, Dean, Settle, Instrumental Methods of Analysis, 7th ed, IBH Book House, New Delhi. 3. Atkins, P.W & Paula, J.D. Physical Chemistry, 10th Ed., Oxford University Press (2014). 4. Kakkar, R. Atomic and Molecular Spectroscopy: Concepts and Applications. Cambridge University Press, 2015. 5. Castellan, G. W. Physical Chemistry 4th Ed., Narosa (2004).6. Banwell, C. N. & McCash, E. M. Fundamentals of Molecular Spectroscopy 4th Ed. Tata McGraw- Hill: New Delhi (2006). 7. Smith, B.C. Infrared Spectral Interpretations:A Systematic Approach. CRC Press, 1998. 8. Moore, W.J., Physical Chemistry Orient Blackswan, 1999.

CEMHTDSE-3B	**Inorganic materials of industrial importance**	**4 credits**

Silicate Industries

1. Glass:

Glassy state and its properties, classification (silicate and non-silicate glasses). Manufacture and processing of glass. Composition and properties of the following types of glasses: Soda lime glass, lead glass, armoured glass, safety glass, borosilicate glass, fluorosilicate, coloured glass, photosensitive glass.

2. Ceramics:

Important clays and feldspar, ceramic, their types and manufacture. High technology ceramics and their applications, superconducting and semiconducting oxides, fullerenes carbon nanotubes and carbon fibre.

3. Cements:

Classification of cement, ingredients and their role, Manufacture of cement and the setting process, quick setting cements.

4. Fertilizers

Different types of fertilizers. Manufacture of the following fertilizers: Urea, ammonium nitrate, calcium ammonium nitrate, ammonium phosphates; polyphosphate, superphosphate, compound and mixed fertilizers, potassium chloride, potassium sulphate.

5. Surface Coatings

Objectives of coatings surfaces, preliminary treatment of surface, classification of surface coatings. Paints and pigments-formulation, composition and related properties. Oil paint, Vehicle, modified oils, Pigments, toners and lakes pigments, Fillers, Thinners, Enamels, emulsifying agents. Special paints (Heat retardant, Fire retardant, Eco-friendly paint, Plastic paint), Dyes, Wax polishing, Water and Oil paints, additives, Metallic coatings (electrolytic and electroless), metal spraying and anodizing.

6.Batteries

Primary and secondary batteries, battery components and their role, Characteristics of Battery. Working of following batteries: Pb acid, Li-Battery, Solid state electrolyte battery. Fuel cells, Solar cell and polymer cell.

7.Alloys

Classification of alloys, ferrous and non-ferrous alloys, Specific properties of elements in alloys. Manufacture of Steel (removal of silicon decarbonization, demanganization, desulphurization dephosphorisation) and surface treatment (Ar and heat treatment, nitriding, carburizing). Composition and properties of different types of steels.

8.Catalysis

General principles and properties of catalysts, homogenous catalysis (catalytic steps and examples) and heterogenous catalysis (catalytic steps and examples) and their industrial applications, Deactivation or regeneration of catalysts. Phase transfer catalysts, application of zeolites as catalysts.

9.Chemical explosives

Origin of explosive properties in organic compounds, preparation and explosive properties of lead azide, PETN, cyclonite (RDX). Introduction to rocket propellants.

Reference Book

1.E. Stocchi: Industrial Chemistry, Vol-I, Ellis Horwood Ltd. UK. 2. R. M. Felder, R. W. Rousseau: Elementary Principles of Chemical Processes, Wiley Publishers, New Delhi. 3. W. D. Kingery, H. K. Bowen, D. R. Uhlmann: Introduction to Ceramics, Wiley Publishers, New Delhi. 4. J. A. Kent: Riegel's Handbook of Industrial Chemistry, CBS Publishers, New Delhi. 5. P. C. Jain, M. Jain: Engineering Chemistry, Dhanpat Rai & Sons, Delhi. 6. R. Gopalan, D.

Venkappayya, S. Nagarajan: Engineering Chemistry, Vikas Publications, New Delhi. 7. Sharma, B.K. & Gaur, H. Industrial Chemistry, Goel Publishing House, Meerut (1996).	

CEMHPDSE-3A	**Instrumental methods of Chemical Analysis**	**2 credits**

1) Safety Practices in the Chemistry Laboratory
2) Determination of Cobalt and Nickel from mixture
3) Study of Electronic Transitions in Organic Molecules (i.e., acetone in water)
4) IR Absorption Spectra (Study of Aldehydes and Ketones)
5) Determination of Calcium, Iron, and Copper in Food by Atomic Absorption
6) Potentiometric Titration of a Chloride - Iodide Mixture
7) Analysis of illicit drugs.
8) Detection in the field and confirmation in the laboratory of flammable accelerants or explosives
9) Detection of steroids.
10) Detection of pollutants from wastes.
11) Fibre analysis
12) Titration curve of an amino acid.
13) Determination of the void volume of a gel filtration column.
14) Quantitative Analysis of Mixtures by Gas Chromatography (i.e.,chloroform and carbon tetrachloride)
15) Separation of Carbohydrates by HPLC
16) Determination of Caffeine in Beverages by HPLC
17) Cyclic Voltammetry of the Ferrocyanide/ Ferricyanide Couple
18) Nuclear Magnetic Resonance
19) Use of fluorescence to do “presumptive tests” to identify blood or other body fluids.
20) Use of “presumptive tests” for anthrax or cocaine
21) Collection, preservation, and control of blood evidence being used for DNA testing
22) Use of capillary electrophoresis with laser fluorescence detection for nuclear DNA (Y chromosome only or multiple chromosome)
23) Use of sequencing for the analysis of mitochondrial DNA

Reference Books

1.Skoog, D.A. Holler F.J. & Nieman, T.A. Principles of Instrumental Analysis, Cengage Learning India Ed. 2. Willard, H.H., Merritt, L.L., Dean, J. & Settoe, F.A. Instrumental Methods of Analysis, 7th Ed. Wadsworth Publishing Company Ltd., Belmont, California, USA, 1988.

CEMHPDSE-3B	**Inorganic materials of industrial importance**	**2 credits**

List of Practicals

1. Determination of free acidity in ammonium sulphate fertilizer.
2. Estimation of Calcium in Calcium ammonium nitrate fertilizer.
3. Estimation of phosphoric acid in superphosphate fertilizer.
4. Electroless metallic coatings on ceramic and plastic material.
5. Determination of composition of dolomite (by complexometric titration).
6. Analysis of (Cu, Ni); (Cu, Zn) in alloy or synthetic samples.
7. Analysis of Cement. 8. Preparation of pigment (zinc oxide).

Reference Books

1 E. Stocchi: Industrial Chemistry, Vol-I, Ellis Horwood Ltd. UK. 2. R. M. Felder, R. W. Rousseau: Elementary Principles of Chemical Processes, Wiley Publishers, New Delhi. 3. W. D. Kingery, H. K. Bowen, D. R. Uhlmann: Introduction to Ceramics, Wiley Publishers, New Delhi. 4. J. A. Kent: Riegel's Handbook of Industrial Chemistry, CBS Publishers, New Delhi. 5.P. C. Jain, M. Jain: Engineering Chemistry, Dhanpat Rai & Sons, Delhi. 6. R. Gopalan, D. Venkappayya, S. Nagarajan: Engineering Chemistry, Vikas Publications, New Delhi. 7. Sharma, B.K. & Gaur, H. Industrial Chemistry, Goel Publishing House, Meerut (1996).

CEMHTDSE- 4	**Green Chemistry**	**4 Credit**

Theory: Introduction to Green Chemistry, Principles of Green Chemistry and Designing a Chemical synthesis, Examples, Future Trends

1. **Introduction to Green Chemistry: (4L)**
   What is Green Chemistry? Need for Green Chemistry. Goals of Green Chemistry. Limitations/ Obstacles in the pursuit of the goals of Green Chemistry
2. **Principles of Green Chemistry and Designing a Chemical synthesis: (27 L)**
   Twelve principles of Green Chemistry with their explanations and examples and special emphasis on the following:
   Designing a Green Synthesis using these principles; Prevention of Waste/ byproducts; maximum incorporation of the materials used in the process into the final products, Atom Economy, calculation of atom economy of the rearrangement, addition, substitution and elimination reactions.
   Prevention/ minimization of hazardous/ toxic products reducing toxicity. risk = (function) hazard × exposure; waste or pollution prevention hierarchy.
   Green solvents– supercritical fluids, water as a solvent for organic reactions, ionic liquids, fluorous biphasic solvent, PEG, solventless processes, immobilized solvents and how to compare greenness of

solvents.

Energy requirements for reactions – alternative sources of energy: use of microwaves and ultrasonic energy.

Selection of starting materials; avoidance of unnecessary derivatization – careful use of blocking/protecting groups.

Use of catalytic reagents (wherever possible) in preference to stoichiometric reagents; catalysis and green chemistry, comparison of heterogeneous and homogeneous catalysis, biocatalysis, asymmetric catalysis and photocatalysis.

Prevention of chemical accidents designing greener processes, inherent safer design, principle of ISD "What you don't have cannot harm you", greener alternative to Bhopal Gas Tragedy (safer route to carcarbaryl) and Flixiborough accident (safer route to cyclohexanol) subdivision of ISD, minimization, simplification, substitution, moderation and limitation. Strengthening/ development of analytical techniques to prevent and minimize the generation of hazardous substances in chemical processes.

3. Examples of Green Synthesis/ Reactions and some real world cases

(25 L)

Green Synthesis of the following compounds: adipic acid, catechol, disodium iminodiacetate (alternative to Strecker synthesis)

Microwave assisted reactions in water: Hofmann Elimination, methyl benzoate to benzoic acid, oxidation of toluene and alcohols; microwave assisted reactions in organic solvents Diels-Alder reaction and Decarboxylation reaction

Ultrasound assisted reactions: sonochemical Simmons-Smith Reaction (Ultrasonic alternative to Iodine)

Surfactants for carbon dioxide – replacing smog producing and ozone depleting solvents with CO_2 for precision cleaning and dry cleaning of garments.

Designing of Environmentally safe marine antifoulant.

Right fit pigment: synthetic azopigments to replace toxic organic and inorganic pigments.

An efficient, green synthesis of a compostable and widely applicable plastic (poly lactic acid) made from corn.

Healthier Fats and oil by Green Chemistry: Enzymatic Inter esterification for production of no Trans-Fats and Oils

Development of Fully Recyclable Carpet: Cradle to Cradle Carpeting

4. Future Trends in Green Chemistry: (4L)

Oxidation reagents and catalysts; Biomimetic multifunctional reagents; Combinatorial green chemistry; Proliferation of solventless reactions; co crystal controlled solid state synthesis (C_2S_3); Green chemistry in sustainable development.

Reference Books

1. Anastas, P.T. & Warner, J.K.: Green Chemistry - Theory and Practical, Oxford University Press (1998). 2. Matlack, A.S. Introduction to Green Chemistry, Marcel Dekker (2001). 3. Cann, M.C. & Connely, M.E. Real-World cases in Green Chemistry, American Chemical Society, Washington (2000). 4. Ryan, M.A. & Tinnesand, M. Introduction to Green Chemistry, American Chemical Society, Washington (2002). 5. Lancaster, M. Green Chemistry: An Introductory Text RSC Publishing, 2nd Edition, 2010.

CEMHPDSE- 4	**Green Chemistry**	**2 Credit**

Practical

1. **Safer starting materials:**
   Preparation and characterization of nanoparticles of gold using tea leaves.

2. **Using renewable resources:**
   Preparation of biodiesel from vegetable/ waste cooking oil.

3. **Avoiding waste:**
   Principle of atom economy.
   Use of molecular model kit to stimulate the reaction to investigate how the atom economy can illustrate Green Chemistry.

   Preparation of propene by two methods can be studied

   a. Triethylamine ion + OH^- → propene + trimethylpropene + water

   b. 1-propanol $\xrightarrow{H_2SO_4/\Delta}$ Propene + water

   Other types of reactions, like addition, elimination, substitution and rearrangement should also be studied for the calculation of atom economy.

4. **Use of enzymes as catalysts:**
   Benzoin condensation using Thiamine Hydrochloride as a catalyst instead of cyanide.

5. **Alternative Green solvents:**
   Extraction of D-limonene from orange peel using liquid CO_2 prepared from dry ice. Mechanochemical solvent free synthesis of azomethines

6. **Alternative sources of energy:**
   Solvent free, microwave assisted one pot synthesis of phthalocyanine complex of copper (II).
   Photoreduction of benzophenone to benzopinacol in the presence of sunlight.

Reference Books

1.Anastas, P.T & Warner, J.C. Green Chemistry: Theory and Practice, Oxford University Press (1998). 2. Kirchoff, M. & Ryan, M.A. Greener approaches to undergraduate chemistry experiment. American Chemical Society, Washington DC (2002). 3. Ryan, M. A. Introduction to Green Chemistry, Tinnesand; (Ed), American Chemical Society, Washington DC (2002). 4. Sharma, R. K.; Sidhwani, I. T. & Chaudhari, M. K. I. K. Green Chemistry Experiment: A monograph International Publishing House Pvt Ltd. New Delhi. Bangalore CISBN 978-93-81141-55-7 (2013). 5. Cann, M. C. & Connelly,M. E. Real world cases in Green Chemistry, American Chemical Society (2008). 6. Cann, M. C. & Thomas, P. Real world cases in Green Chemistry, American Chemical Society (2008). 7. Lancaster, M. Green Chemistry: An Introductory Text RSC Publishing, 2nd Edition, 2010. 8. Pavia, D. L., Lampman, G. M., Kriz, G. S. & Engel, R. G. Introduction to Organic Laboratory Techniques: A Microscale and Macro Scale Approach, W.B.Saunders, 1995.

DSEPROJ-5	**Project**	**6 credits**

CEMHSE-2A	**Pharmaceutical Chemistry**	**2 Credit**

1. **Drugs & Pharmaceuticals: (16L)**
   Drug discovery, design and development; Basic Retrosynthetic approach. Synthesis of the representative drugs of the following classes: analgesics agents, antipyretic agents, anti-inflammatory agents (Aspirin, paracetamol, lbuprofen); antibiotics (Chloramphenicol); antibacterial and antifungal agents (Sulphonamides; Sulphanethoxazol, Sulphacetamide, Trimethoprim); antiviral agents (Acyclovir), Central Nervous System agents (Phenobarbital, Diazepam), Cardiovascular (Glyceryl trinitrate), antilaprosy (Dapsone), HIV-AIDS related drugs (AZT- Zidovudine).

2. **Fermentation: (7L)**
   Aerobic and anaerobic fermentation. Production of (i) Ethyl alcohol and citric acid, (ii) Antibiotics; Penicillin, Cephalosporin, Chloromycetin and Streptomycin, (iii) Lysine, Glutamic acid, Vitamin B2, Vitamin B12 and Vitamin C.

3. **Hands on Practical: (7L)**
   Preparation of Aspirin and its analysis. Preparation of magnesium bisilicate (Antacid).

Reference Books

1. Patrick, G. L. Introduction to Medicinal Chemistry, Oxford University Press, UK, 2013.2.Singh, H. & Kapoor, V.K. Medicinal and Pharmaceutical Chemistry, Vallabh Prakashan, Pitampura, New Delhi, 2012. 3. Foye, W.O., Lemke, T.L. & William, D.A.: Principles of Medicinal Chemistry 4th edition B.I Waverly Pvt. Ltd. New Delhi

CEMHSE-2B	Analytical Clinical Biochemistry	2 Credit

Concept of carbohydrates, Proteins, Enzyms, Lipids; Biochemistry of disease

1 Review of concepts of core course (8L)

Carbohydrates: Biological importance of carbohydrates, metabolism cellular currency of energy(ATP), Glycolysis, Alcoholic and lactic acid fermentation, Kreb's cycle. Isolation and characterization of polysaccharides.

Proteins: Classification, biological importance, Primary and secondary and tertiary structures of proteins: α-helix β –pleated sheets, isolation, characterization , denaturation of proteins.

Enzymes: Nomenclature, characterization (mention of ribozymes) and classification, active site, Mechanism of enzyme action, stereospecificity of enzymes, coenzymes and cofactors, enzyme inhibitors, Introduction to biocatalysts, Importance in "Green Chemistry" and chemical industry.

Lipids: Classification, Biological importance of triglycerides and phosphoglycerides and cholesterol: Lipid membrane, Liposomes and their biological functions and underlying applications, Lipoproteins, Properties, functions and biochemical functions of steroid hormones, Biochemistry of peptide hormones.

2. Biochemistry of disease : A diagnostic approach by blood / urine analysis (12L)

Blood: Composition and functions of blood, blood coagulation ,blood collection and preservation of samples. Anaemia, Regulation, estimation and interpretation of data for blood sugar, urea, creatinine, cholesterol and bilirubin

Urine: Sampling and preservation, composition and estimation of constituents of normal and pathological urine.

3. Hands on Practical (10L)

Identification and estimation of the following :

i)Carbohydrates- qualitative and quantitative
ii) Lipids- qualitative
iii) Determination of iodine number of oil
iv) Determination of saponification number of oil
v) Determination of cholesterol using Liebermann-Burchard reaction.
vi) Proteins- qualitative
vii) Isolation of protein

viii) Determination of protein by the Biuret reaction ix) Determination of nucleic acids **Reference Books** 1.Cooper, T.G. Tool of Biochemistry. Wiley-Blackwell (1977). 2. Wilson, K. & Walker, J. Practical Biochemistry. Cambridge University Press (2009). 3. Varley, H., Gowenlock, A.H & Bell, M.: Practical Clinical Biochemistry, Heinemann, London (1980). 4. Devlin, T.M., Textbook of Biochemistry with Clinical Correlations, John Wiley & Sons, 2010. 5. Berg, J.M., Tymoczko, J.L. & Stryer, L. Biochemistry, W.H. Freeman, 2002. 6. Talwar, G.P. & Srivastava, M. Textbook of Biochemistry and Human Biology, 3rd Ed. PHI Learning. 7. Nelson, D.L. & Cox, M.M. Lehninger Principles of Biochemistry, W.H. Freeman, 2013. 8. O. Mikes, R.A. Chalmers: Laboratory Handbook of Chromatographic Methods, D. Van Nostrand & Co., 1961.	

N.B.: Please ignore typographical error if any.